

## Chapter 9 Reaction rates

Time permitted: 30 minutes

	Section	Number of questions	Marks available
<b>A</b>	Multiple choice	15	15
<b>B</b>	Short answer	5	15
	<b>Total</b>		<b>30</b>

Scale:

<b>A+</b>	29–30	<b>A</b>	26–28	<b>B</b>	23–25	<b>C</b>	19–22	<b>D</b>	15–18	<b>E</b>	9–14	<b>UG</b>	0–8
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### Section A Multiple choice (15 marks)

Section A consists of 15 questions, each worth one mark. Each question has only one correct answer. Circle the correct answer. Attempt all questions. Marks will not be deducted for incorrect answers. You are advised to spend no more than 15 minutes on this section.

- 1 Why does reaction rate increase as temperature increases?
- ☒ **A** Particles move faster and this increases the probability of collision.
  - ☐ **B** There are more particles at higher temperatures so more of them collide.
  - ☐ **C** The activation energy is lowered.
  - ☐ **D** Particles require a set temperature in which to begin colliding.

For questions 2 and 3 refer to the chemical reaction below.



- 2 Which does not occur during the reaction?
- ☒ **A** The concentration of  $\text{H}_2\text{O}$  increases.
  - ☐ **B** The mass of  $\text{CaCO}_3$  decreases.
  - ☐ **C** The volume of  $\text{CO}_2$  increases.
  - ☐ **D** The concentration of  $\text{Ca}^{2+}$  increases.
- 3 The rate of the reaction could not be determined by:
- ☐ **A** recording the change in mass of the  $\text{CaCO}_3$  over time.
  - ☐ **B** recording the mass of  $\text{CO}_2$  lost over time.
  - ☒ **C** measuring the final mass of  $\text{CaCO}_3$ .
  - ☐ **D** collecting and recording the volume of  $\text{CO}_2$  gas produced over time.

# Lesson 36 Reaction Rates Answers Key

**Judd E. Hollander**



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