

# Annotated Solution

## 2018 USNCO Local Exam

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### 1 Solutions

1. We can use Avogadro's constant, and then do some dimensional analysis.

$$225 \text{ g O}_2 = \frac{1 \text{ mol O}_2}{32.0 \text{ g O}_2} \times \frac{2 \text{ mol O}}{1 \text{ mol O}_2} = \frac{6.92 \times 10^{23} \text{ atoms}}{\text{mol O}} = \boxed{6.42 \times 10^{24} \text{ atoms}}$$

Thus, the answer is **C**.

2. Because the same number of moles must remain before and after dilution, we can set up the following equation:

$$\begin{aligned} M_1 V_1 &= M_2 V_2 \\ (1.50 \text{ M})(V_1) &= (0.000 \text{ M})(0.500 \text{ L}) \\ (1.50 \text{ M})(V_1) &= 0.150 \text{ M} \cdot \text{L} \\ V_1 &= 0.100 \text{ L} \end{aligned}$$

Thus is **100 mL**, or **A**.

3. We first need to write the reaction in order to perform the stoichiometry. The reaction is as follows:



Now we can do the stoichiometry:

$$10.00 \text{ g Cu} = \frac{1 \text{ mol Cu}}{63.55 \text{ g Cu}} \times \frac{1 \text{ mol H}_2\text{O}}{2 \text{ mol Cu}} \times \frac{18.02 \text{ g H}_2\text{O}}{1 \text{ mol H}_2\text{O}} = \boxed{1.417 \text{ g H}_2\text{O}}$$

Thus, the answer is **B**.

4. Set a standard amount of sample at an arbitrary 100.0 g. Thus, there are 48.41 g Cu, 19.02 g As, and 32.57 g S. We can divide by the molar masses of each of the compounds:

$$\begin{aligned} 48.41 \text{ g Cu} &= \frac{1 \text{ mol Cu}}{63.55 \text{ g Cu}} = 0.7618 \text{ mol} \\ 19.02 \text{ g As} &= \frac{1 \text{ mol As}}{74.92 \text{ g As}} = 0.2539 \text{ mol} \\ 32.57 \text{ g S} &= \frac{1 \text{ mol S}}{32.06 \text{ g S}} = 1.016 \text{ mol} \end{aligned}$$

Dividing by the lowest number, there there will be the ratio of Cu : As : S = 3 : 1 : 4. Thus, the formula is **Cu<sub>3</sub>AsS<sub>4</sub>**, or **C**.

# National Standardized Chemistry Exam

**Ying-Ying Zheng**



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W. Blood-Ryan,1912      **The Chemical News and Journal of Industrial Science** ,1928      California State Journal of Medicine ,1905      **“The” Athenaeum** ,1873      **Program-based Review and Assessment** Martha L. A. Stassen,2004  
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### **Table of Contents National Standardized Chemistry Exam**

1. Understanding the eBook National Standardized Chemistry Exam
  - The Rise of Digital Reading National Standardized Chemistry Exam
  - Advantages of eBooks Over Traditional Books
2. Identifying National Standardized Chemistry Exam
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an National Standardized Chemistry Exam
  - User-Friendly Interface
4. Exploring eBook Recommendations from National Standardized Chemistry Exam
  - Personalized Recommendations
  - National Standardized Chemistry Exam User Reviews and Ratings
  - National Standardized Chemistry Exam and Bestseller Lists

5. Accessing National Standardized Chemistry Exam Free and Paid eBooks
  - National Standardized Chemistry Exam Public Domain eBooks
  - National Standardized Chemistry Exam eBook Subscription Services
  - National Standardized Chemistry Exam Budget-Friendly Options
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  - ePub, PDF, MOBI, and More
  - National Standardized Chemistry Exam Compatibility with Devices
  - National Standardized Chemistry Exam Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of National Standardized Chemistry Exam
  - Highlighting and Note-Taking National Standardized Chemistry Exam
  - Interactive Elements National Standardized Chemistry Exam
8. Staying Engaged with National Standardized Chemistry Exam
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers National Standardized Chemistry Exam
9. Balancing eBooks and Physical Books National Standardized Chemistry Exam
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection National Standardized Chemistry Exam
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine National Standardized Chemistry Exam
  - Setting Reading Goals National Standardized Chemistry Exam
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of National Standardized Chemistry Exam
  - Fact-Checking eBook Content of National Standardized Chemistry Exam
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

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- Interactive and Gamified eBooks

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