

## Stoichiometry: Multiple Choice

1. The limiting reactant of a chemical reaction is

- A) the reactant of a chemical process that is not consumed completely
- B) the product of a chemical process that is consumed completely
- C) the product of a chemical process that is not consumed completely
- D) the reactant of a chemical process that is consumed completely
- E) none of the above

ANS: D

1. Consider the following reaction;  $\text{Ca}(\text{OH})_{2(aq)} + \text{Al}_2(\text{SO}_4)_{3(aq)} \rightarrow \text{CaSO}_{4(s)} + \text{Al}(\text{OH})_{3(s)}$

When balanced, the mole ratio for calcium hydroxide reacting with aluminum sulphate is

- A) 1:1
- B) 2:1
- C) 2:3
- D) 1:2
- E) 3:1

ANS: E

1. Sodium phosphate,  $\text{Na}_3\text{PO}_4$ , can be prepared by combining sodium hydroxide and hydrogen phosphate as shown in the following balanced chemical equation:



If 2.5 mol of  $\text{H}_2\text{PO}_4$  are reacted, the number of moles of water that will be produced is

- A) 1.0
- B) 2.5
- C) 3.0
- D) 5.0
- E) 7.5

ANS: E

1. Potassium chloride and oxygen gas are formed upon heating potassium chlorate. The following balanced equation shows this decomposition process:



The number of moles of  $\text{KClO}_3$  required to produce 5.0 mol of oxygen is

- A) 1.0
- B) 2.0
- C) 3.0
- D) 3.3
- E) 5.0

ANS: D

1. Determine the number of grams of  $\text{CO}_2$  that can be produced from 48.0 g of the hydrocarbon  $\text{C}_3\text{H}_8$ , according to the following equation:



- A) 37.6 g
- B) 50.2 g
- C) 87.8 g
- D) 151 g
- E) 188.4 g

1. The limiting reactant in a chemical reaction is the \_\_\_\_\_ that is \_\_\_\_\_.

- A) Reactant, in excess
- B) Product, left over
- C) Reactant, used up
- D) Product, used up

ANS: C

# Multiple Choice Questions On Stoichiometry

**Terry C. Jones**



## **Multiple Choice Questions On Stoichiometry:**

## Multiple Choice Questions On Stoichiometry Book Review: Unveiling the Magic of Language

In an electronic era where connections and knowledge reign supreme, the enchanting power of language has become more apparent than ever. Its capability to stir emotions, provoke thought, and instigate transformation is actually remarkable. This extraordinary book, aptly titled "**Multiple Choice Questions On Stoichiometry**," compiled by a very acclaimed author, immerses readers in a captivating exploration of the significance of language and its profound effect on our existence. Throughout this critique, we will delve to the book's central themes, evaluate its unique writing style, and assess its overall influence on its readership.

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