

Momentum Study Guide

What is Momentum?

- Momentum is the quantity of motion of a moving object.
- It is a vector quantity, meaning it has both magnitude and direction.
- The formula for momentum: $p=m \cdot v$
- where:
 p = momentum (kg·m/s)
 m = mass (kg)
 v = velocity (m/s)

Conservation of Momentum

- In a closed system, the total momentum before a collision or interaction is equal to the total momentum after. $p_{\text{total}} = p_{\text{total}}$

Study Guide: Understanding Acceleration

What is Acceleration?

- Acceleration is the rate at which an object's velocity changes over time.
- It is a vector quantity, meaning it has both **magnitude** and **direction**.
- Formula of acceleration

$$a = \frac{\Delta v}{\Delta t}$$

Where:

- a = acceleration (m/s²)
- Δv = change in velocity (m/s)
 ($v_{\text{final}} - v_{\text{initial}}$)
- Δt = time interval (s)

When Does Acceleration Happen?

- It occurs when there is a change in:
 - **Speed** (increasing or decreasing).
 - **Direction** (turning or curving).
 - **Both speed and direction** simultaneously.
- **When an Object Speeds Up**
 - This is called **positive acceleration**.
 - Example: A car increasing its speed from 0 m/s to 30 m/s.
- **When an Object Slows Down**
 - This is called **negative acceleration** or **deceleration**.

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web we then changed the volume between 5ml and 20ml and recorded the pressure at each this part of the lab showed that volume and pressure are inversely proportionate as the pressure would decrease if the volume was increased part b of the experiment looked at the relationship between temperature and pressure

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web procedure 1 the relationship between pressure and volume set the plunger of the syringe to 1 10 the capacity of the syringe the 2 0 ml mark if using a 20 ml syringe the volume is read from the edge of the bottom rib of the plunger attach the syringe to the connector at the end of sensor

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web he concluded that at a constant pressure the volume of the gas would expand at a particular rate for each degree of temperature rise that being 1 267 per degree in 1874 victor regnault obtained better experimental results showing this number to be 1 273 and concluded that the pressure would approach zero at 1 273 degrees c

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web the volume of 1 00 mol 1 00 mol of any gas at stp standard temperature 273 15 k and pressure 1 atm is measured to be
22 414 l 22 414 l we can substitute 101 325 kpa 101 325 kpa for pressure 22 414 l 22 414 l for volume and 273 15 k 273 15 k
for temperature into the ideal gas equation and solve for r

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web chemists sometimes make comparisons against a standard temperature and pressure stp for reporting properties of
gases 273 15 k and 1 atm 101 325 kpa 1 at stp one mole of an ideal gas has a volume of about 22 4 l this is referred to as the
standard molar volume figure 9 18

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web this lab explores the relationships between pressure volume and temperature in an ideal gas the ideal gas law $pV = nRT$ is
used and with it we can explore it by seeing how the pressure and volumes changed in the tubes each time that we

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web figure pageindex 10 since the number of moles in a given volume of gas varies with pressure and temperature changes
chemists use standard temperature and pressure 273 15 k and 1 atm or 101 325 kpa to report properties of gases

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introduction the ideal gas law is a commonly used formula in chemistry it relates pressure temperature and volume this
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web volume pressure kpa linear pressure kpa volume ml pressure kpa in part b we observed the relationship between

temperature and pressure the pressure of the 125 ml flask was at 101 36 kpa when the temperature was at 23 2 c

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web pressure increases even though there were errors in the experiment the goal of the experiment was met the goal of declaring what type of relationship it is was clear using the data pressure kpa temperature c temperature k constant k p t 95 82 kpa 16 0 c 289 k 0 33 97 90 kpa 19 0 c 292 k 0 34 98 97 kpa 99 22 kpa 101 4 kpa 19 1 c

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