

Ions

How are ions made from neutral atoms?

Why?

You have learned that not all atoms of an element are the same. Variation in the number of neutrons results in different isotopes of the element. In this activity we will explore another variation that can take place—the loss and gain of electrons. The exchange of electrons between atoms is a very common way for chemical change to take place. We will see it many times throughout the year.

1. Use Model 1 to complete the following table.

	Metal or Nonmetal	Is the number of protons the same in the atom and the ion?	Is the number of neutrons the same in the atom and the ion?	Is the number of electrons the same in the atom and the ion?	Charge on the ion
Lithium	metal	yes	yes	no	1+
Magnesium	metal	yes	yes	no	2+
Aluminum	metal	yes	yes	no	3+
Fluorine	nonmet	yes	yes	no	1-
Oxygen	nonmetal	yes	yes	no	2-
Nitrogen	nonmet	yes	yes	no	3-



2. Based on the table you completed in Question 1, what distinguishes a neutral atom from an ion?

e^-

3. Examine the isotope symbols in Model 1.

a. Where is the ion charge located in the isotope symbol?

upper right ${}^3\text{Li}^{1+}$



b. Is a charge indicated on the neutral atoms? If yes, where is it located?

NO



4. Which subatomic particle carries a positive charge?

proton +

5. Which subatomic particle carries a negative charge?

electron -

6. Propose a mathematical equation to calculate the charge on an ion from the number of protons and electrons in an ion. Confirm that your equation works using two positive ion examples and two negative ion examples from Model 1.

$$\boxed{\#p - \#e = \text{charge}}$$

$\text{Li} \quad 3 - 2 = +1$

$\text{Mg} \quad 12 - 10 = +2$

$\text{F} \quad 9 - 10 = -1$

$\text{N} \quad 7 - 10 = -3$

Ions Packet Answers

Peter J. Mikulecky, Chris Hren



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