

BOARD QUESTION PAPER: MARCH 2022

CHEMISTRY

Time: 3 Hrs

Max. Marks: 70

General Instructions:

The question paper is divided into four sections.

- (1) **Section A:** Q. No. 1 contains Ten multiple choice type of questions carrying One mark each.
Q. No. 2 contains Eight very short answer type of questions carrying One mark each.
- (2) **Section B:** Q. No. 3 to Q. No. 14 are Twelve short answer type of questions carrying Two marks each. (Attempt any Eight).
- (3) **Section C:** Q. No. 15 to Q. No. 26 are Twelve short answer type of questions carrying Three marks each. (Attempt any Eight).
- (4) **Section D:** Q. No. 27 to Q. No. 31 are Five long answer type of questions carrying Four marks each. (Attempt any Three).
- (5) Use of log table is allowed. Use of calculator is not allowed.
- (6) Figures to the right indicate full marks.
- (7) For each multiple choice type of question, it is mandatory to write the correct answer along with its alphabet, e.g. (a)...../(b)...../(c)...../(d)..... etc.

No marks(s) shall be given, if ONLY the correct answer or the alphabet of the correct answer is written. Only the first attempt will be considered for evaluation.

SECTION – A

Q.1. Select and write the correct answer for the following multiple choice type of questions:

[10]

- (i) The co-ordination number of atoms in body centred cubic structure (bcc) is _____.
(a) 4 (b) 6 (c) 8 (d) 12
- (ii) In calculating osmotic pressure, the concentration of solute is expressed in _____.
(a) molarity (b) molality
(c) mole fraction (d) percentage mass
- (iii) The enthalpy change for the chemical reaction $\text{H}_2\text{O}_{(s)} \longrightarrow \text{H}_2\text{O}_{(g)}$ is called enthalpy of _____.
(a) vapourisation (b) fusion
(c) combustion (d) sublimation
- (iv) Which of the following transition element shows maximum oxidation state?
(a) Sc (b) Fe (c) Mn (d) V
- (v) The correct formula for the complex compound, sodium hexacyanoferrate (III) is _____.
(a) $\text{Na} [\text{Fe}(\text{CN})_6]$ (b) $\text{Na}_2 [\text{Fe}(\text{CN})_6]$
(c) $\text{Na}_3 [\text{Fe}(\text{CN})_6]$ (d) $\text{Na}_4 [\text{Fe}(\text{CN})_6]$

Mcomst Year Question Paper

RJ Shavelson



Mcomst Year Question Paper:

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Table of Contents Mcomst Year Question Paper

1. Understanding the eBook Mcomst Year Question Paper
 - The Rise of Digital Reading Mcomst Year Question Paper
 - Advantages of eBooks Over Traditional Books
2. Identifying Mcomst Year Question Paper
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Mcomst Year Question Paper
 - User-Friendly Interface
4. Exploring eBook Recommendations from Mcomst Year Question Paper
 - Personalized Recommendations
 - Mcomst Year Question Paper User Reviews and Ratings
 - Mcomst Year Question Paper and Bestseller Lists

5. Accessing Mcomst Year Question Paper Free and Paid eBooks
 - Mcomst Year Question Paper Public Domain eBooks
 - Mcomst Year Question Paper eBook Subscription Services
 - Mcomst Year Question Paper Budget-Friendly Options
6. Navigating Mcomst Year Question Paper eBook Formats
 - ePub, PDF, MOBI, and More
 - Mcomst Year Question Paper Compatibility with Devices
 - Mcomst Year Question Paper Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Mcomst Year Question Paper
 - Highlighting and Note-Taking Mcomst Year Question Paper
 - Interactive Elements Mcomst Year Question Paper
8. Staying Engaged with Mcomst Year Question Paper
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Mcomst Year Question Paper
9. Balancing eBooks and Physical Books Mcomst Year Question Paper
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Mcomst Year Question Paper
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Mcomst Year Question Paper
 - Setting Reading Goals Mcomst Year Question Paper
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Mcomst Year Question Paper
 - Fact-Checking eBook Content of Mcomst Year Question Paper
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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