

### Objective type

- Q.1 Which of the following is not a general characteristic of equilibria involving physical processes?
- (1) Equilibrium is possible only in a closed system at a given temperature.
  - (2) All measurable properties of the system remain constant.
  - (3) All the physical processes stop at equilibrium.
  - (4) The opposing processes occur at the same rate and there is dynamic but stable condition.
- Q.2 Which of the following statements is incorrect?
- (1) In equilibrium mixture of ice and water kept in perfectly insulated flask mass of ice and water does not change with time.
  - (2) The intensity of red colour increases when oxalic acid is added to a solution containing iron (III) nitrate and potassium thiocyanate.
  - (3) On addition of catalyst the equilibrium constant value is not affected.
  - (4) Equilibrium constant for a reaction with negative  $\Delta H$  value decreases as the temperature increases.
- Q.3 When hydrochloric acid is added to cobalt nitrate solution at room temperature, the following reaction takes place and the reaction mixture becomes blue. On cooling the mixture it becomes pink. On the basis of this information mark the correct answer.
- $$\underset{\text{(pink)}}{[\text{Co}(\text{H}_2\text{O})_6]^{3+}}(\text{aq}) + 4\text{Cl}^{-}(\text{aq}) \rightleftharpoons \underset{\text{(blue)}}{[\text{CoCl}_4]^{2-}}(\text{aq}) + 6\text{H}_2\text{O}(\text{l})$$
- (1)  $\Delta H > 0$  for the reaction
  - (2)  $\Delta H < 0$  for the reaction
  - (3)  $\Delta H = 0$  for the reaction
  - (4) The sign of  $\Delta H$  cannot be predicted on the basis of this information.
- Q.4 What will be the correct order of vapour pressure of water, acetone and ether at  $30^{\circ}\text{C}$ . Given that among these compounds, water has maximum boiling point and ether has minimum boiling point?
- (1) Water < ether < acetone
  - (2) Water < acetone < ether
  - (3) Ether < acetone < water
  - (4) Acetone < ether < water
- Q.5 In which of the following reactions, the equilibrium remains unaffected on addition of small amount of argon at constant volume?
- (1)  $\text{H}_2(\text{g}) + \text{I}_2(\text{g}) \rightleftharpoons 2\text{HI}(\text{g})$
  - (2)  $\text{PCl}_5(\text{g}) \rightleftharpoons \text{PCl}_3(\text{g}) + \text{Cl}_2(\text{g})$
  - (3)  $\text{N}_2(\text{g}) + 3\text{H}_2(\text{g}) \rightleftharpoons 2\text{NH}_3(\text{g})$
  - (4) The equilibrium will remain unaffected in all the three cases

### Assertion & Reason

- Q.6 Assertion (A): In the dissociation of  $\text{PCl}_5$  at constant pressure and temperature addition of helium at equilibrium increases the dissociation of  $\text{PCl}_5$ .
- Reason (R) : Helium removes  $\text{Cl}_2$  from the field of action.
- (1) Both A and R are true and R is correct explanation of A.
  - (2) Both A and R are true but R is not correct explanation of A.
  - (3) A is true but R is false.
  - (4) Both A and R are false.

# Mcqs Of Environmental Chemistry

**Lei Huang**



## **Mcqs Of Environmental Chemistry:**

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In a time characterized by interconnectedness and an insatiable thirst for knowledge, the captivating potential of verbal expression has emerged as a formidable force. Its ability to evoke sentiments, stimulate introspection, and incite profound transformations is genuinely awe-inspiring. Within the pages of "**Mcqs Of Environmental Chemistry**," a mesmerizing literary creation penned by a celebrated wordsmith, readers set about an enlightening odyssey, unraveling the intricate significance of language and its enduring effect on our lives. In this appraisal, we shall explore the book's central themes, evaluate its distinctive writing style, and gauge its pervasive influence on the hearts and minds of its readership.

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