

CK-12 Chemistry - Basic  
Answer Key**Chapter 8: Ionic and Metallic Bonding**

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**8.1 Ions****Practice 8.1**

Write the ground state configuration for the metal calcium, and predict the ion form it must become to be isoelectric with the nearest noble gas.

Answer:

Ca:  $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2$  or  $[\text{Ar}]4s^2$

The nearest noble gas to calcium is  $[\text{Ar}]$ . So calcium will typically lose two electrons to become isoelectric with  $[\text{Ar}]$ .

**Practice 8.2**

Write the ground state configuration for the nonmetal sulfur, and predict the ion form it must become to be isoelectric with the nearest noble gas.

Answer:

S:  $1s^2 2s^2 2p^6 3s^2 3p^4$  or  $[\text{Ne}]3s^2 3p^4$

The nearest noble gas to sulfur is  $[\text{Ar}]$ , so sulfur will typically gain two electrons to become  $[\text{Ne}] 3s^2 3p^6$  which is isoelectric with  $[\text{Ar}]$

**Lesson Review Questions***Questions*

1. Draw electron dot diagrams for one metal and one nonmetal from the periodic table.
2. Predict which of the following will likely become a cation or an anion.
  - a. Ca
  - b. Na
  - c. F
  - d. Br
  - e. S
3. Write the ground state configuration for the following elements and predict the ion form it must become to be isoelectric with the nearest noble gas.
  - a. Be

# Ionic And Metallic Bonding Answer Key

**John T. Moore**



## **Ionic And Metallic Bonding Answer Key:**

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