

Page 3	Mark Scheme	Syllabus	Paper
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- 2 (a) (i) positive and negative ions [1]
regular pattern / opposite charges closer than the same charge [1]
- (ii) so that charges cancel / ions may not have the same charge [1]
- (iii) Any three of:
high melting point or boiling point
hard
brittle
soluble in water / insoluble in organic solvents
conduct (electricity) in liquid state or in aqueous solution / non-conductors or poor conductor (when solid) [3]
- (b) correct formula [1]
correct charges [1]
6x and 2x around oxygen [1]

[Total: 8]

- 3 (a) (i) roast or heat or burn in air / roast or heat or burn in oxygen [1]
need both of the above
- (ii) $\text{ZnO} + \text{C} \rightarrow \text{Zn} + \text{CO}$ / $2\text{ZnO} + \text{C} \rightarrow 2\text{Zn} + \text{CO}_2$ / $\text{ZnO} + \text{CO} \rightarrow \text{Zn} + \text{CO}_2$ [1]
- (b) (i) $\text{ZnO} + \text{H}_2\text{SO}_4 \rightarrow \text{ZnSO}_4 + \text{H}_2\text{O}$ [1]
- (ii) zinc reduces / gives electrons / displaces (copper / cobalt / nickel ions) [1]
forming copper / cobalt / nickel (metal which is precipitated) [1]
- (c) (i) $\text{Zn}^{2+} + 2\text{e}^- \rightarrow \text{Zn}$ [1]
- (ii) $\text{OH}^- \rightarrow \text{H}_2\text{O} + \text{O}_2 + \dots\text{e}^-$ (1) only
 $4\text{OH}^- \rightarrow 2\text{H}_2\text{O} + \text{O}_2 + 4\text{e}^-$ [2]
- (iii) sulfuric acid / hydrogen sulfate [1]
ACCEPT: sulfuric acid

Mark Scheme For 20 N 13

**I. A. Mekeel, Charles Esterly
Severn, Stephen B. Hopkins**

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