

WORKSHEET – Naming Compounds KEY

Part I – Naming Compounds and Numbering Atoms

FORMULA	NAME	# of DIFFERENT ATOMS	TOTAL # OF ATOMS
1. CaCl_2	Calcium chloride	1 Ca; 2 Cl	3
2. AgCl	<i>Silver chloride</i>	1 Ag; 1 Cl	2
3. MgO	<i>Magnesium oxide</i>	1 Mg; 1 O	2
4. NaBr	<i>Sodium bromide</i>	1 Na; 1 Br	2
5. Al_2O_3	<i>Aluminum oxide</i>	2 Al; 3 O	5
6. KI	<i>Potassium iodide</i>	1 K; 1 I	2
7. ZnCl_2	<i>Zinc chloride</i>	1 Zn; 2 Cl	3
8. Li_2O	<i>Lithium oxide</i>	2 Li; 1 O	3
9. BaO	<i>Barium oxide</i>	1 Ba; 1 O	2
10. CaBr_2	<i>Calcium bromide</i>	1 Ca; 2 Br	3
11. MgCl_2	<i>Magnesium chloride</i>	1 Mg; 2 Cl	3
12. AgI	<i>Silver iodide</i>	1 Ag; 1 I	2
13. ZnS	<i>Zinc sulphide</i>	1 Zn; 1 S	2
14. BaF_2	<i>Barium fluoride</i>	1 Ba; 2 F	3
15. Ca_3P_2	<i>Calcium phosphide</i>	3 Ca; 2 P	5
16. Na_2O	<i>Sodium oxide</i>	2 Na; 1 O	3
17. AlN	<i>Aluminum nitride</i>	1 Al; 1 N	2
18. K_3N	<i>Potassium nitride</i>	3 K; 1 N	4
19. Ca_2C	<i>Calcium carbide</i>	2 Ca; 1 C	3
20. K_2O	<i>Potassium oxide</i>	2 K; 1 O	3

Naming Ionic Compounds Practice Answer Key

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