#### AP Chemistry Chapter 14 Answers - Zumdahl

#### 14.27

a.  $HClO_4(aq) + H_2O(1) \rightarrow H_3O^*(aq) + ClO_4^-(aq)$ . Only the forward reaction is indicated since  $HClO_4$  is a strong acid and is basically 100% dissociated in water. For acids, the dissociation reaction is commonly written without water as a reactant. The common abbreviation for this reaction is:  $HClO_4(aq) \rightarrow H^*(aq) + ClO_4^-(aq)$ . This reaction is also called the  $K_a$  reaction as the equilibrium constant for the reaction is called  $K_a$ .

b. Propanioc acid is a weak acid, so it is only partially dissociated in water. The dissociation reaction is:  $CH_3CH_2CO_2H(aq) + H_2O(l) \Rightarrow H_3O^* + CH_3CH_2CO_2(aq)$  or  $CH_3CH_2CO_2H(aq) \Rightarrow$  $H^*(aq) + CH_3CH_2CO^*(aq)$ .

NH<sub>4</sub>\* is a weak acid. Similar to propanoic acid, the dissociation reaction is:

$$NH_4^*(aq) + H_2O(1) \Rightarrow H_3O^*(aq) + NH_3(aq) \text{ or } NH_4^*(aq) \Rightarrow H^*(aq) + NH_3(aq)$$

#### 14.29

An acid is a proton (H\*) donor and a base is a proton acceptor. A conjugate acid-base pair differs by only a proton (H\*).

	Acid	Base	Conjugate Base of Acid	Conjugate Acid of Base
28.	H <sub>2</sub> CO <sub>3</sub>	H <sub>2</sub> O	HCO <sub>3</sub>	H <sub>7</sub> O*
Ib.	C <sub>2</sub> H <sub>2</sub> NH <sup>4</sup>	H <sub>2</sub> O	C <sub>5</sub> H <sub>5</sub> N	H <sub>2</sub> O*
C.	$C_4H_5NH^*$	HCO <sub>3</sub>	C <sub>5</sub> H <sub>5</sub> N	$H_2CO_3$

#### 14.31

Strong acids have a K<sub>a</sub> >> 1 and weak acid have K<sub>a</sub> < 1. Table 14.2 in the text lists some K<sub>a</sub> values for weak acids. K<sub>a</sub> values for strong acids are hard to determine so they are not listed in the text. However, there are only a few common strong acids so if you memorize the strong acids, then all other acids will be weak acids. The strong acids to memorize are HCl, HBr, HI, HNO<sub>3</sub>, HClO<sub>4</sub>, and H<sub>2</sub>SO<sub>4</sub>.

- a. HClO<sub>4</sub> is a strong acid.
- HOCl is a weak acid (K<sub>a</sub> = 3.5 x 10<sup>-8</sup>).
- e. H<sub>2</sub>SO<sub>4</sub> is a strong acid.
- d. H<sub>2</sub>SO<sub>3</sub> is a weak diprotic acid with K<sub>a1</sub> and K<sub>a2</sub> values less than one.

#### 14.33

The K<sub>a</sub> value is directly related to acid strength. As K<sub>a</sub> increases, acid strength increases. For water, use K<sub>w</sub> when comparing the acid strength of water to other species. The K<sub>a</sub> values are:

HClO<sub>4</sub>: strong acid (
$$K_a >> 1$$
); HClO<sub>2</sub>:  $K_a = 1.2 \times 10^{-2}$   
NH<sub>4</sub>\*:  $K_a = 5.6 \times 10^{-10}$ ; H<sub>2</sub>O:  $K_a = K_w = 1.0 \times 10^{-14}$ 

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**David Baud** 

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