

AP Chemistry Chapter 14 Answers – Zumdahl

14.27

a.
 $\text{HClO}_4(\text{aq}) + \text{H}_2\text{O}(\text{l}) \rightarrow \text{H}_3\text{O}^+(\text{aq}) + \text{ClO}_4^-(\text{aq})$. Only the forward reaction is indicated since HClO_4 is a strong acid and is basically 100% dissociated in water. For acids, the dissociation reaction is commonly written without water as a reactant. The common abbreviation for this reaction is: $\text{HClO}_4(\text{aq}) \rightarrow \text{H}^+(\text{aq}) + \text{ClO}_4^-(\text{aq})$. This reaction is also called the K_a reaction as the equilibrium constant for the reaction is called K_a .

b.
 Propanoic acid is a weak acid, so it is only partially dissociated in water. The dissociation reaction is: $\text{CH}_3\text{CH}_2\text{CO}_2\text{H}(\text{aq}) + \text{H}_2\text{O}(\text{l}) \rightleftharpoons \text{H}_3\text{O}^+ + \text{CH}_3\text{CH}_2\text{CO}_2^-(\text{aq})$ or $\text{CH}_3\text{CH}_2\text{CO}_2\text{H}(\text{aq}) \rightleftharpoons \text{H}^+(\text{aq}) + \text{CH}_3\text{CH}_2\text{CO}_2^-(\text{aq})$.

c.
 NH_4^+ is a weak acid. Similar to propanoic acid, the dissociation reaction is:



14.29

An acid is a proton (H^+) donor and a base is a proton acceptor. A conjugate acid-base pair differs by only a proton (H^+).

	Acid	Base	Conjugate Base of Acid	Conjugate Acid of Base
a.	H_2CO_3	H_2O	HCO_3^-	H_3O^+
b.	$\text{C}_5\text{H}_5\text{NH}^+$	H_2O	$\text{C}_5\text{H}_5\text{N}$	H_3O^+
c.	$\text{C}_6\text{H}_5\text{NH}^+$	HCO_3^-	$\text{C}_6\text{H}_5\text{N}$	H_2CO_3

14.31

Strong acids have a $K_a \gg 1$ and weak acids have $K_a < 1$. Table 14.2 in the text lists some K_a values for weak acids. K_a values for strong acids are hard to determine so they are not listed in the text. However, there are only a few common strong acids so if you memorize the strong acids, then all other acids will be weak acids. The strong acids to memorize are HCl , HBr , HI , HNO_3 , HClO_4 , and H_2SO_4 .

- HClO_4 is a strong acid.
- HOCl is a weak acid ($K_a = 3.5 \times 10^{-8}$).
- H_2SO_4 is a strong acid.
- H_2SO_3 is a weak diprotic acid with K_{a1} and K_{a2} values less than one.

14.33

The K_a value is directly related to acid strength. As K_a increases, acid strength increases. For water, use K_w when comparing the acid strength of water to other species. The K_a values are:



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1. Understanding the eBook Modern Chemistry Answer Key Review
 - The Rise of Digital Reading Modern Chemistry Answer Key Review
 - Advantages of eBooks Over Traditional Books
2. Identifying Modern Chemistry Answer Key Review
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in a Modern Chemistry Answer Key Review
 - User-Friendly Interface
4. Exploring eBook Recommendations from Modern Chemistry Answer Key Review
 - Personalized Recommendations
 - Modern Chemistry Answer Key Review User Reviews and Ratings
 - Modern Chemistry Answer Key Review and Bestseller Lists
5. Accessing Modern Chemistry Answer Key Review Free and Paid eBooks
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 - Modern Chemistry Answer Key Review Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Modern Chemistry Answer Key Review
 - Highlighting and Note-Taking Modern Chemistry Answer Key Review
 - Interactive Elements Modern Chemistry Answer Key Review
8. Staying Engaged with Modern Chemistry Answer Key Review

- Joining Online Reading Communities
- Participating in Virtual Book Clubs
- Following Authors and Publishers Modern Chemistry Answer Key Review
- 9. Balancing eBooks and Physical Books Modern Chemistry Answer Key Review
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Modern Chemistry Answer Key Review
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Modern Chemistry Answer Key Review
 - Setting Reading Goals Modern Chemistry Answer Key Review
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Modern Chemistry Answer Key Review
 - Fact-Checking eBook Content of Modern Chemistry Answer Key Review
 - Distinguishing Credible Sources
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 - Utilizing eBooks for Skill Development
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