

Chapter 6 Practice Test: Chemical Bonding

In the space provided, write the letter of the term or phrase that best completes each statement or best answers each question.

- _____ 1. The charge on an ion is
a. always positive.
b. always negative.
c. either positive or negative.
d. zero.
- _____ 2. According to the octet rule, a calcium atom has a tendency to
a. lose one electron.
b. lose two electrons.
c. gain one electron.
d. gain two electrons.
- _____ 3. If a compound forms by ionic bonding, which is *not* true?
a. A positively charged atom or group of atoms attracts a negatively charged atom or group of atoms.
b. The net charge of the compound is zero.
c. The compound contains just two atoms, each of opposite charge.
d. Several ions group together in a tightly packed structure.
- _____ 4. The only property listed that is *not* characteristic of ionic compounds is
a. high melting point.
b. hardness.
c. lack of crystal structure.
d. good conductor of electricity.
- _____ 5. Which formula listed below represents a polyatomic ion?
a. NH_4^+
b. H_2SO_4
c. NaCl
d. H_2O
- _____ 6. The melting points of ionic compounds are higher than the melting points of molecular compounds because
a. ionic substances tend to vaporize at room temperature.
b. ionic substances are brittle.
c. attractive forces between ions are greater than the attractive forces between molecules.
d. the numbers of positive and negative charges are equal in an ionic compound.

Modern Chemistry Chapter 4 2 Review Answers

**Holt Rinehart & Winston, Holt,
Rinehart and Winston Staff**



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Modern Chemistry Holt Rinehart & Winston, Holt, Rinehart and Winston Staff, 2001 Organic Synthesis Michael Smith, 2011-07-12 A reactions oriented course is a staple of most graduate organic programs and synthesis is taught either as a part of that course or as a special topic Ideally the incoming student is an organic major who has a good working knowledge of basic reactions stereochemistry and conformational principles In fact however many often most of the students in a first year graduate level organic course have deficiencies in their undergraduate work are not organic majors and are not synthetically inclined To save students much time catching up this text provides a reliable and readily available source for background material that will enable all graduate students to reach the same high level of proficiency in organic chemistry Produced over many years with extensive feedback from students taking an organic chemistry course this book provides a reaction based approach The first two chapters provide an introduction to functional groups these are followed by chapters reviewing basic organic transformations e g oxidation reduction The book then looks at carbon carbon bond formation reactions and ways to disconnect a bigger molecule into simpler building blocks Most chapters include an extensive list of questions to test the reader's understanding There is also a new chapter outlining full retrosynthetic analyses of complex molecules which highlights common problems made by scientists The book is intended for graduate and postgraduate students scientific researchers in chemistry New publisher new edition extensively updated and corrected Over 950 new references with more than 6100 references in total Over 600 new reactions and figures replaced or updated Over 300 new homework problems from the current literature to provide nearly 800 problems to test reader understanding of the key principles **Principles of Modern Chemistry** David W. Oxtoby, H. Pat Gillis, Laurie J. Butler, 2016-01-01 Long considered the standard for honors and high level mainstream general chemistry courses PRINCIPLES OF MODERN CHEMISTRY continues to set the standard as the most modern rigorous and chemically and mathematically accurate text on the market This authoritative text features an atoms first approach and thoroughly revised chapters on Quantum Mechanics and Molecular Structure Chapter 6 Electrochemistry Chapter 17 and Molecular Spectroscopy and Photochemistry Chapter 20 In addition the text utilizes mathematically accurate and artistic atomic and molecular orbital art and is student friendly without compromising its rigor End of chapter study aids focus on only the most important key objectives equations and concepts making it easier for students to locate chapter content while applications to a wide range of disciplines such as biology chemical engineering biochemistry and medicine deepen students understanding of the relevance of chemistry beyond the classroom **Volume 1: Modern Electrochemistry** John O'M. Bockris, Amulya K.N. Reddy, 2007-05-08 This book had its nucleus in some lectures given by one of us J O M B in a course on electrochemistry to students of energy conversion at the University of Pennsylvania It was there that he met a number of people trained in chemistry physics biology metallurgy and materials science all of whom wanted to know something about electrochemistry The concept of writing a

book about electrochemistry which could be understood by people with very varied backgrounds was thereby engendered. The lectures were recorded and written up by Dr Klaus Muller as a 293 page manuscript. At a later stage A K N R joined the effort it was decided to make a fresh start and to write a much more comprehensive text. Of methods for direct energy conversion the electrochemical one is the most advanced and seems the most likely to become of considerable practical importance. Thus conversion to electrochemically powered transportation systems appears to be an important step by means of which the difficulties of air pollution and the effects of an increasing concentration in the atmosphere of carbon dioxide may be met. Corrosion is recognized as having an electrochemical basis. The synthesis of nylon now contains an important electrochemical stage. Some central biological mechanisms have been shown to take place by means of electrochemical reactions. A number of American organizations have recently recommended greatly increased activity in training and research in electrochemistry at universities in the United States.

Independent and Weekly Review, 1914 Solutions Guide for Introductory Chemistry James F. Hall, 1996 **The Electrical Review**, 1904 *Modern Chemistry* Nicholas D. Tzimopoulos, 1990 *The Modern Review* Ramananda Chatterjee, 1920. Includes section Reviews and notices of books.

Lab Experiments for Modern Chemistry Tzimopoulos, 1990 **Solutions Guide, Introductory Chemistry, a Foundation, Introductory Chemistry, Basic Chemistry, Fourth Edition, Zumdahl** James F. Hall, 2000 **Basic Concepts of Chemistry, Study Guide** Leo J. Malone, 2000. The text's three main goals are to introduce chemistry as a living relevant science to encourage learning and critical thinking and to help readers overcome the math difficulties that impede their progress in chemistry. Designed to help readers master the principles of general chemistry. As a prep book it promotes active involvement with the material. There are special features throughout that reinforce concepts and help to develop strong problem solving and study skills. Updated to include an Interactive Learning Ware problems CD containing several of the chapter ending problems from the book in an interactive tutorial with feedback to help readers set up and solve problems.

World History, 2000 India Rubber World and Electrical Trades Review, 1927 **The Chautauquan**, 1885
Physics Art Hobson, 2007. For a one semester course in liberal arts physics Hobson has four unifying themes. How do we know the significance of post Newtonian physics, modern physics, energy and the social context of physics. These themes become evident in the writing and pedagogy throughout the fourth edition. **Chemist and Druggist**, 1927 **The United States Catalog**, 1921 **The United States Catalog Supplement, January 1918-June 1921** Eleanor E. Hawkins, 1921 The United States Catalog; Books in Print January 1, 1912 Marion Effie Potter, 1921

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