

AP Chemistry Pacing Guide

Unit 4: Chemical Reactions - Introduction for Reactions - Net Ionic Equations - Representations of Reactions - Physical and Chemical Changes - Stoichiometry - Introduction to Titration - Types of Chemical Reactions - Introduction to Acid-Base Reactions - Oxidation-Reduction (Redox) Reactions	BIG IDEA 1: Scale, Proportion, and Quantity BIG IDEA 3: Transformations	- I can identify evidence of chemical and physical changes in matter. - I can represent changes in matter with a balanced chemical or net ionic equation: a. For physical changes; b. For given information about the identity of the reactants and/or product. - For ions in a given chemical reaction I can represent a given chemical reaction or physical process with a consistent particulate model. - I can explain the relationship between macroscopic characteristics and bond interactions for: a. Chemical processes; b. Physical processes. - I can explain changes in the amounts of reactants and products based on the balanced reaction equation for a chemical process. - I can identify the equivalence point in a titration based on the amounts of the titrant and analyte, assuming the titration reaction goes to completion. - I can identify a reaction as acid-base, oxidation-reduction, or precipitation. - I can identify species as Brønsted-Lowry acids, bases, and/or conjugate acid-base pairs, based on proton-transfer involving those species. - I can represent a balanced redox reaction equation using half-reactions.	10 days
Unit 5: Kinetics - Reaction Rates - Introduction to Rate Law - Concentration Changes Over Time - Elementary Reactions - Collision Model - Reaction Energy Profile - Introduction to Reaction Mechanism - Mechanism Reaction Mechanism and Rate Law - Steady-State	BIG IDEA 3: Transformations BIG IDEA 6: Energy	- I can explain the relationship between the rate of a chemical reaction and experimental parameters. - I can represent experimental data with a consistent rate law expression. - I can identify the rate law expression of a chemical reaction using data that show how the concentrations of reaction species change over time. - I can represent an elementary reaction as a rate law expression using stoichiometry. - I can explain the relationship between the rate of an elementary reaction and the frequency, energy, and orientation of molecular collisions. - I can represent the activation energy and overall energy change in an elementary reaction using a reaction energy profile. - I can identify the components of a reaction mechanism. - I can identify the rate law for a reaction from a mechanism in which the first step is rate limiting. - I can identify the rate law for a reaction from a mechanism in which the	10 days

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