

### balance nuclear reaction equations

aluminum-26 undergoes positron emission  
potassium-40 undergoes beta decay (electron)  
uranium-238 undergoes alpha decay



### electron configurations

Write the electron configuration for C.

Write the electron configuration for Cr.

Write the shorthand notation for Bi.



### NOMENCLATURE [naming compounds, writing formulas]

Write the formula for each of the following compounds:

- a. Sulfur hexafluoride  $\text{SF}_6$
- b. Lithium nitride  $\text{Li}_3\text{N}$
- c. Chromium (III) carbonate  $\text{Cr}_2(\text{CO}_3)_3$
- d. Tin (II) chloride  $\text{SnCl}_2$
- e. Ammonium acetate  $\text{NH}_4\text{C}_2\text{H}_3\text{O}_2$  (or  $\text{CH}_3\text{COONH}_4$ )
- f. Mercury (I) chloride  $\text{Hg}_2\text{Cl}_2$
- g. Potassium bromate  $\text{KBrO}_3$
- h. Hydrosulfuric acid  $\text{H}_2\text{S(aq)}$
- i. Chloric acid  $\text{HClO}_3$
- j. Sulfurous acid  $\text{H}_2\text{SO}_3$

Name each of the following compounds.

- a.  $\text{CuSO}_4$  Copper (II) sulfate
- b.  $\text{AlF}_3$  aluminum fluoride
- c.  $\text{HI}$  hydrogen iodide
- d.  $\text{NO}$  nitrogen monoxide
- e.  $\text{H}_2\text{Se(aq)}$  hydroselenic acid
- f.  $\text{HNO}_3$  nitric acid
- g.  $\text{HNO}_2$  nitrous acid
- h.  $\text{NaHSO}_4$  sodium hydrogen sulfate

### Half-life

Tritium ( $\text{H-3}$ ) is a radioactive isotope of hydrogen with a half-life of 12.3 years. How long would it take for a 40.0 g sample to decay down to 1.25 g?

$$40 \xrightarrow{1} 20 \xrightarrow{2} 10 \xrightarrow{3} 5 \xrightarrow{4} 2.5 \xrightarrow{5} 1.25$$

Fe-61 has a half-life of 6.00 min. Of a 100.0 mg sample, how much will remain after 18.0 min?

$$3 \text{ half lives} \quad 100 \xrightarrow{1} 50 \xrightarrow{2} 25 \xrightarrow{3} 12.5 \text{ mg}$$

### Average atomic mass

The element copper has naturally occurring isotopes with mass numbers of 63 and 65.

The relative abundance and atomic masses are 69.2% for a mass of 63amu and 30.8% for a mass of 65amu. Calculate the average atomic mass of copper.

$$\begin{array}{rcl} 63 \times .692 & = & 43.596 \rightarrow 44 \\ 65 \times .308 & = & 20.02 \rightarrow 20 \\ + & & \\ \hline & & 63.62 \rightarrow 64 \end{array}$$

# Modern Chemistry Review Answers Chapter 11

**Morris Hein, Susan Arena, Cary Willard**



## Modern Chemistry Review Answers Chapter 11:

*Modern Chemistry* Holt Rinehart & Winston, Holt, Rinehart and Winston Staff, 2001      **Solutions Guide for Introductory Chemistry** James F. Hall, 1996      **Foundations of College Chemistry, Student Solutions Manual** Morris Hein, Susan Arena, Cary Willard, 2022-03-09 Check your work and better retain the lessons of Foundations of College Chemistry The Foundations of College Chemistry 16th Edition Student Solutions Manual is an essential study resource and aid that allows users of the widely read text to check their work against the detailed solutions contained in this book This Solutions Manual unlocks the full value of the practice problem and exercise focused 16th edition of Foundations of College Chemistry Students will be able to study efficiently and effectively and learn from common errors with the additional guidance offered by this book      *Solutions Guide, Introductory Chemistry, a Foundation, Introductory Chemistry, Basic Chemistry, Fourth Edition, Zumdahl* James F. Hall, 2000      *Saturday Review*, 1863      **Independent and Weekly Review**, 1914      **Book Review Digest**, 1916      **Instructor's Manual and Test Bank to Accompany Basic Concepts of Chemistry** Malone, 2000-04-18      **The Saturday Review of Politics, Literature, Science and Art**, 1901

**Foundations of Chemistry** Morris Hein, Susan Arena, 1996 A textbook introducing matter atomic theory ionization and other aspects of chemistry to the high school student      **General Chemistry** Ralph H. Petrucci, William S. Harwood, 1993 General Chemistry Principles and Modern Applications is recognized for its superior problems lucid writing and precision of argument This updated and expanded edition retains the popular and innovative features of previous editions including Feature Problems follow up Integrative and Practice Exercises to accompany every in chapter Example and Focus On application boxes as well as new Keep in Mind marginal notes Topics covered include atoms and the atomic theory chemical compounds and reactions gases Thermochemistry electrons in atoms chemical bonding liquids solids and intermolecular forces chemical kinetics principles of chemical equilibrium acids and bases electrochemistry representative and transitional elements and nuclear and organic chemistry For individuals interested in a broad overview of chemical principles and applications      **Modern Electroplating** Electrochemical Society, 1963      *The Saturday Review of Politics, Literature, Science, Art, and Finance*, 1863      **Introductory Chemistry** Steven S. Zumdahl, 1993      *Basic Chemistry* Steven S. Zumdahl, 1996      **High School Biology Review** Kim Magloire, Princeton Review (Firm), 1998-03 If trudging through your textbook to study and complete homework assignments has become a frustrating grind then get ready for a smooth ride to higher test scores and outstanding grades with The Princeton Review's High School Biology Review We tell it to you straight thoroughly explaining the important topics you'll need to understand to prepare for quizzes and tests complete homework assignments effectively and earn higher grades We've carefully examined biology textbooks just like yours to make sure that this book includes all the material essential to a thorough review In this guide we cover The Chemistry of Life Life Functions Biodiversity The Cell Reproduction Ecology Cellular Respiration Genetics Laboratory Skills Photosynthesis Modern Genetics

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