

## Chemistry Student Edition - Basic Answer Key

### Chapter 10: The Mole

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#### 10.1 The Mole Concept

##### Questions

1. Imagine that you have 1 mole of coins, each of which is 1.5 mm thick. If they were placed in a single stack, how tall would the stack be (in km)? If the closest distance between the earth and the moon is 356,400 km, would the coins reach the moon? (Hint: Use the technique of dimensional analysis.)
2. Calculate the number of moles of N that must be present in 3 moles of  $\text{NH}_3$ .
3. Calculate the number of moles of oxygen that must be present in 2 moles of  $\text{C}_6\text{H}_{12}\text{O}_6$ .
4. How many moles of gold are present in a sample containing  $1.81 \times 10^{24}$  gold atoms?
5. How many atoms are in 0.065 moles of Hg?
6. Calculate the number of  $\text{Br}^-$  ions in 0.0038 moles of  $\text{CaBr}_2$ .
7. Calculate the molar mass of  $\text{K}_2\text{Cr}_2\text{O}_7$ .
8. Calculate the molar mass of dinitrogen tetroxide ( $\text{N}_2\text{O}_4$ ).

##### Answers

1.  $6.02 \times 10^{23} \text{ coins} \times \frac{1.5 \text{ mm}}{\text{coin}} \times \frac{1 \text{ m}}{1 \times 10^3 \text{ mm}} \times \frac{1 \text{ km}}{1 \times 10^3 \text{ m}} = 6.02 \times 10^{17} \text{ km}$   
Yes, the coins would reach the moon.
2. There is one mole N/mole  $\text{NH}_3$ , so three moles  $\text{NH}_3$  contain three moles N
3. 8 moles O/mole glucose, so there are  $8 \times 2 = 12$  moles O in two moles glucose.
4.  $\frac{1.81 \times 10^{24} \text{ atoms}}{6.02 \times 10^{23} \text{ atoms / mole}} = 3 \text{ moles}$
5.  $(0.065 \text{ moles})(6.02 \times 10^{23} \text{ atoms/mole}) = 1.06 \times 10^{21} \text{ atoms}$
6. There are 2 Br ions/molecule of  $\text{CaBr}_2$   
 $\frac{2 \text{ ions}}{\text{molecule}} \times \frac{6.02 \times 10^{23} \text{ molecules}}{\text{mole}} \times 3.8 \times 10^{-3} \text{ moles} = 4.57 \times 10^{21} \text{ ions}$
7.  $\text{K } 2 \times 39.1 = 78.2$   
 $\text{Cr } 2 \times 52.0 = 104.0$   
 $\text{O } 7 \times 16 = 112.0$   
molar mass of  $\text{K}_2\text{Cr}_2\text{O}_7 = 294.2 \text{ grams/mole}$
8.  $\text{N } 2 \times 14 = 28.0$   
 $\text{O } 4 \times 16 = 64.0$   
molar mass of  $\text{K}_2\text{Cr}_2\text{O}_7 = 92.0 \text{ grams/mole}$

# Modern Chemistry Answer Key Chapter 10

**Ying Liu**



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