Molarity Practice

Calculate for the unknown variables. M =n/v

- 1. Explain what molarity/concentration means?
- What is the molarity of a solution that contains 5.50 moles of calcium chloride in 2.5 liters of solution?

TeachSimple

- 3. What is the molarity of a solution that contains 3.39 x 10⁻² moles hydrobromic acid in 1 liter of solution?
- 4. A beaker contains 250g C₆H₁₂O₆ (glucose) in 1,500 mL of solution. What is the molarity?
- A flask contains 325g grams iron II oxide and has a molarity of 2M. What is the volume of the solution?
- 6. Calculate the moles if a flask contains 3.25M potassium nitrate in 750 mL of solution.
- 7. A cup contains 85.31 grams ammonium acetate in 3.27 liters of solution calculate the molarity.
- 8. What is the molarity of a solution that contains 10.5 grams magnesium oxide in 250 ml. of solution?
- 9. What mass of potassium phosphate would you use to make 2.5 liter of 1.25 M solution?

Molarity Practice Chemistry Unit Lesson

Mike Jess

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