

Molarity Calculations

Calculate the molarities of the following solutions:



- 1) 2.3 moles of sodium chloride in 0.45 liters of solution.

$$M = \frac{2.3 \text{ moles}}{0.45 \text{ L}} = 5.111 \frac{\text{mol}}{\text{L}} = \boxed{5.11 \text{ M}}$$



- 2) 1.2 moles of calcium carbonate in 1.22 liters of solution.

$$M = \frac{1.2 \text{ mol}}{1.22 \text{ L}} = 0.983606557 \frac{\text{mol}}{\text{L}} = \boxed{0.98 \text{ M}}$$



- 3) 0.09 moles of sodium sulfate in 12 mL of solution.

12 mL = 0.012 L

$$M = \frac{0.09 \text{ mol}}{0.012 \text{ L}} = \boxed{7.5 \text{ M}}$$



- 4) 0.75 moles of lithium fluoride in 65 mL of solution.

65 mL = 0.065 L

$$M = \frac{0.75 \text{ mol}}{0.065 \text{ L}} = 11.53846154 \text{ M} = \boxed{11.54 \text{ M}}$$



- 5) 0.8 moles of magnesium acetate in 5 liters of solution.

$$M = \frac{0.8 \text{ mol}}{5 \text{ L}} = \boxed{0.16 \text{ M}}$$



- 6) 120 grams of calcium nitrite in 240 mL of solution.

240 mL = 0.240 L

$$n = 120 \text{ g} \left( \frac{1 \text{ mol}}{132.09 \text{ g}} \right)$$

MW<sub>Ca(NO<sub>2</sub>)<sub>2</sub></sub> = 132.09 g/mol

n = 0.91 mol

$$M = \frac{0.91 \text{ mol}}{0.240 \text{ L}} = \boxed{3.79 \text{ M}}$$



- 7) 98 grams of sodium hydroxide in 2.2 liters of solution.

MW<sub>NaOH</sub> = 39.997 g/mol

$$n = 98 \text{ g} \left( \frac{1 \text{ mol}}{39.997 \text{ g}} \right) = 2.45 \text{ mol}$$

$$M = \frac{2.45 \text{ mol}}{2.2 \text{ L}} = \boxed{1.11 \text{ M}}$$



- 8) 1.2 grams of hydrochloric acid in 25 mL of solution.

25 mL = 0.025 L

MW<sub>HCl</sub> = 36.461 g/mol

$$n = (1.2 \text{ g}) \left( \frac{1 \text{ mol}}{36.461 \text{ g}} \right) = 0.033 \text{ mol}$$

$$M = \frac{0.033 \text{ mol}}{0.025 \text{ L}} = \boxed{1.32 \text{ M}}$$



- 9) 45 grams of ammonia in 0.75 L of solution.

MW<sub>NH<sub>3</sub></sub> = 17.031 g/mol

$$n = 45 \text{ g} \left( \frac{1 \text{ mol}}{17.031 \text{ g}} \right) = 2.64 \text{ mol}$$

$$M = \frac{2.64 \text{ mol}}{0.75 \text{ L}} = \boxed{3.52 \text{ M}}$$

# Molarity Calculations Key Answers Show Work

**Fred M. Dewey**



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## **Table of Contents Molarity Calculations Key Answers Show Work**

1. Understanding the eBook Molarity Calculations Key Answers Show Work
  - The Rise of Digital Reading Molarity Calculations Key Answers Show Work
  - Advantages of eBooks Over Traditional Books
2. Identifying Molarity Calculations Key Answers Show Work
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Molarity Calculations Key Answers Show Work
  - User-Friendly Interface
4. Exploring eBook Recommendations from Molarity Calculations Key Answers Show Work
  - Personalized Recommendations
  - Molarity Calculations Key Answers Show Work User Reviews and Ratings
  - Molarity Calculations Key Answers Show Work and Bestseller Lists
5. Accessing Molarity Calculations Key Answers Show Work Free and Paid eBooks
  - Molarity Calculations Key Answers Show Work Public Domain eBooks
  - Molarity Calculations Key Answers Show Work eBook Subscription Services
  - Molarity Calculations Key Answers Show Work Budget-Friendly Options

6. Navigating Molarity Calculations Key Answers Show Work eBook Formats
  - ePub, PDF, MOBI, and More
  - Molarity Calculations Key Answers Show Work Compatibility with Devices
  - Molarity Calculations Key Answers Show Work Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Molarity Calculations Key Answers Show Work
  - Highlighting and Note-Taking Molarity Calculations Key Answers Show Work
  - Interactive Elements Molarity Calculations Key Answers Show Work
8. Staying Engaged with Molarity Calculations Key Answers Show Work
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Molarity Calculations Key Answers Show Work
9. Balancing eBooks and Physical Books Molarity Calculations Key Answers Show Work
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Molarity Calculations Key Answers Show Work
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Molarity Calculations Key Answers Show Work
  - Setting Reading Goals Molarity Calculations Key Answers Show Work
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Molarity Calculations Key Answers Show Work
  - Fact-Checking eBook Content of Molarity Calculations Key Answers Show Work
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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