

# **Whole Unit Review**

1. Calculate the molar mass of organic molecules:

a)  $\text{HCl}$

b)  $\text{H}_2\text{SO}_4$

c)  $\text{FeCl}_2\text{Cl}_2$

a)  $\text{H}_2\text{O}$

b)  $\text{Cu}_2\text{OH}_2$

c)  $\text{Fe}_2\text{SO}_4$

a)  $\text{H}_2\text{SO}_4$

b)  $\text{Fe}_2\text{SO}_4$

c)  $\text{CH}_3\text{COOH}$

a)  $\text{CO}_2$

b)  $\text{FeCl}_2$

c)  $\text{CH}_3\text{CH}_2\text{CH}_2\text{CH}_3$

a)  $\text{CH}_4$

b)  $\text{FeCl}_3$

2. Calculate the mass of the following:

a) 1.00 moles of  $\text{H}_2\text{SO}_4$

a) 0.0044 moles of  $\text{Fe}_2\text{SO}_4$

b) 1.00 moles of  $\text{H}_2\text{SO}_4$

b) 1.00 moles of  $\text{FeCl}_3$

c) 1.00 moles of  $\text{FeCl}_3$

3. Calculate the number of moles in the following:

a) 17.0 g of  $\text{H}_2\text{SO}_4$

a) 0.125 g of  $\text{H}_2\text{SO}_4$

b) 17.0 g of  $\text{H}_2\text{O}$

b) 0.400 g of  $\text{H}_2\text{O}$

c) 17.0 g of  $\text{Fe}$

4. Calculate the molar mass of organic molecules and then calculate the mass of the following:

a) A 0.001 mol sample of  $\text{H}_2\text{SO}_4$  has a mass of 0.001 g

b) A 0.0044 mol sample of  $\text{Fe}_2\text{SO}_4$  has a mass of 0.0044 g

c) The mass of 0.001 mol of  $\text{FeCl}_3$  is 0.001 g

d) A 0.001 mol sample of  $\text{FeCl}_3$  has a mass of 0.001 g

5. Calculate the number of moles of the following:

a) 17.0 g of  $\text{H}_2\text{SO}_4$

a) 0.0044 g of  $\text{Fe}_2\text{SO}_4$

b) 17.0 g of  $\text{H}_2\text{O}$

b) 0.001 mol of  $\text{FeCl}_3$

c) 17.0 g of  $\text{FeCl}_3$

c) 0.001 mol of  $\text{FeCl}_3$

6. Calculate the volume of the following gases at STP:

a) 0.001 mol of  $\text{H}_2$

a) 0.001 mol of  $\text{H}_2$

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**Jin-Ying Zhang**



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