

1/5/16

Mole Ratio Worksheet

Key

① Given this equation: $\text{N}_2 + 3\text{H}_2 \rightarrow 2\text{NH}_3$, write the following molar ratios:

a) N_2 / H_2 1:3

b) N_2 / NH_3 1:2

c) H_2 / NH_3 3:2

② Given the following equation: $8\text{H}_2 + \text{S}_8 \rightarrow 8\text{H}_2\text{S}$, write the following molar ratios:

a) $\text{H}_2 / \text{H}_2\text{S}$ 8:8 \rightarrow 1:1 (reduced)

b) H_2 / S_8 8:1

c) $\text{H}_2\text{S} / \text{S}_8$ 8:1

③ Answer the following questions for this equation: $2\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$

a) What is the $\text{H}_2 / \text{H}_2\text{O}$ molar ratio? 2:2 \rightarrow 1:1 (reduced)

b) Suppose you had 20 moles of H_2 on hand and plenty of O_2 , how many moles of H_2O could you make? $\frac{20 \text{ mole H}_2}{2 \text{ H}_2} \times \frac{2 \text{ H}_2\text{O}}{1} = \boxed{20 \text{ mole H}_2\text{O}}$

c) What is the $\text{O}_2 / \text{H}_2\text{O}$ molar ratio?

1:2

d) Suppose you had 20 moles of O_2 and enough H_2 , how many moles of H_2O could you make? $\frac{20 \text{ mol O}_2}{1 \text{ O}_2} \times \frac{2 \text{ H}_2\text{O}}{1} = \boxed{40 \text{ mol H}_2\text{O}}$

4) Use this equation: $\text{N}_2 + 3\text{H}_2 \rightarrow 2\text{NH}_3$, for the following problems

a) If you used 1 mole of N_2 , how many moles of NH_3 could be produced? $\boxed{2}$

b) If 10 moles of NH_3 were produced, how many moles of N_2 would be required? $\frac{10 \text{ mol NH}_3}{2 \text{ NH}_3} \times \frac{1 \text{ N}_2}{1} = \boxed{5 \text{ mol N}_2}$

c) If 3.00 moles of H_2 were used, how many moles of NH_3 would be made? $\frac{3 \text{ mol H}_2}{3 \text{ H}_2} \times \frac{2 \text{ NH}_3}{1} = \boxed{2 \text{ mol NH}_3}$

d) If 0.600 moles of NH_3 were produced, how many moles of H_2 are required? $\frac{0.6 \text{ mol NH}_3}{2 \text{ NH}_3} \times \frac{3 \text{ H}_2}{1} = \boxed{0.9 \text{ mol H}_2}$

Mole Ratio Worksheet Chemistry Answers

Research and Education Association

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You Can Do Chemistry Sunil Tanna, 2018-11-19 A comprehensive guide to performing mole and stoichiometric calculations with numerous examples as well as questions and answers Covers calculations relating to solids solutions gases and electrolysis plus as limiting and excess reactants chemical yields atom economy and much more Fully up to date with the last international standards including the revised definition of mole which was agreed on November 16th 2018

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Quick Access Research and Education Association,2009-09-25 REA s Quick Access Study Charts contain all the information students teachers and professionals need in one handy reference They provide quick easy access to important facts The charts contain commonly used math formulas historical facts language conjugations vocabulary and more Great for exams classroom reference or a quick refresher on the subject **Study Guide and Solutions Manual to accompany Basic Concepts of Chemistry, 9th Edition** Leo J. Malone,2011-12-07 This is the Study Guide and Solutions Manual to accompany Malone s Basic Concepts of Chemistry *Chemistry* Ken Hughes,Paul B. Kelter,1998-08

Mole Concepts and Stoichiometry Dharshi Bopegedera,2017 **Bite-Sized Chemistry Calculations** John Lambert,2024-07-10 Bite Sized Chemistry Calculations is a series of books on chemistry calculations aimed at helping students overcome the challenges associated with tackling the various types of calculations encountered in different aspects of chemistry focusing on a few topics at a time to facilitate comprehension Written by an experienced chemistry educator each book in the series has been tailored to fully meet the needs of students at all levels especially those taking college level general chemistry courses as well as those following various O level curricula worldwide This part of the series explores the different types of problems and calculations encountered in mass the mole and stoichiometry including the determination of formulae of ionic compounds relative formula masses mass and percent compositions of compounds all aspects of mole calculations empirical and molecular formulae calculations based on chemical equations limiting reagents gas stoichiometry and percent yield The series is packed with many salient features that are meant to facilitate both teaching and learning Some of these include helpful explanations many examples alternative ways to solve problems plenty of practice questions complete answers an appendices With this book you will be well prepared for your exams and boost your performance

CONTENTS 1 Writing the Formulae of Ionic Compounds 2 Formula Masses 3 Mass and Percent Compositions 4 The Mole and Mass 5 The Mole and Number of Particles 6 The Mole and Concentration 7 The Mole and Molar Volume 8 Empirical and Molecular Formulae 9 Chemical Equations 10 Calculations Based on Chemical Equations 11 Limiting Reagent 12 Gas Volume Stoichiometry 13 Percent Yield Answers to Practice Problems Appendices **Chemistry Equations and Answers** Mark Jackson,2006-02-22 This 6 page study guide contains basic chemistry analysis and concepts designed specifically to aid science students

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