

# GIZMO's - Activity: Moles, Avogadro, and Molar Mass- Answer key 2023

Name: \_\_\_\_\_

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## Activity: Moles, Avogadro, and Molar Mass

### Prior Knowledge Questions

1. In the image, note a dozen eggs, a dozen donuts and a dozen roses. How many of each item do you have? 12 eggs, 12 donuts, and 12 roses.
2. Would a dozen of each object have the same mass? No
3. Suppose you have a dozen silver atoms, a dozen oxygen atoms, and a dozen lithium atoms. Even though you have the same number of each, will they all have the same mass? Explain.  
They will not all have the same mass because although there may be an equal amount of each one, they don't all weigh the same or take up the same amount of space.



### Gizmo Warm-up

On the AVOGADRO CONSTANT tab, place the copper (Cu) atom on the nano-balance on the left, which will show the average atomic mass of copper rather than the mass of a single copper atom.



1. What is the average mass of a copper atom? 63.546 u

The unit "u" refers to **unified atomic mass units**. (Officially, 1 u is one-twelfth the mass of a C-12 atom.)

2. To know how many atoms are in a gram of copper, use the larger balance on the right. Press **Add atoms** to put a scoop of atoms in the weighing dish, and keep adding until the balance registers between 1 and 2 grams. If you don't seem to be making much progress, adjust the exponent using the slider, which will make the scoop size bigger.

How many atoms did you need to add?  $9.6 \times 10^{21}$  atoms of copper

# Moles And Mass Answer Key

**Daniel C. Harris, Chuck Lucy**

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author names.

## Moles And Mass Answer Key:

**Student Solutions Manual** Darrell Ebbing, 2005-03-17 Provides solutions to odd numbered Practice Problems General Problems and Cumulative Skills Problems plus answers to Review Questions *Chemistry Class XI - SBPD Publications* Dr. Subhash Chandra Rastogi, ,Er. Meera Goyal, 2022-02-17 Content 1 Some Basic Concepts of Chemistry 2 Structure of Atom 3 Classification of Elements and Periodicity in Properties 4 Chemical Bonding and Molecular Structure 5 States of Matter 6 Thermodynamics 7 Equilibrium 8 Redox Reactions 9 Hydrogen 10 s Block Elements 11 p Block Elements 12 Organic Chemistry Some Basic Principles and Techniques 13 Hydrocarbons 14 Environmental Chemistry I Appendix II Log antilog Table *GO TO Objective NEET 2021 Chemistry Guide 8th Edition* Disha Experts, **Experimental Organic Chemistry** Daniel R. Palleros, 2000-02-04 This cutting edge lab manual takes a multiscale approach presenting both micro semi micro and macroscale techniques The manual is easy to navigate with all relevant techniques found as they are needed Cutting edge subjects such as HPLC bioorganic chemistry multistep synthesis and more are presented in a clear and engaging fashion **Chemistry Class 11 - [Bihar & JAC]** Dr. S. C. Rastogi, ,Er. Meera Goyal, 2022-07-20 Syllabus Unit I Some Basic Concepts of Chemistry Unit II Structure of Atom Unit III Classification of Elements and Periodicity in Properties Unit IV Chemical Bonding and Molecular Structure Unit V States of Matter Gases and Liquids Unit VI Chemical Thermodynamics Unit VII Equilibrium Unit VIII Redox Reactions Unit IX Hydrogen Unit X s Block Elements Alkali and Alkaline earth metals Group 1 and Group 2 Elements Unit XI Some p Block Elements General Introduction to p Block Elements Unit XII Organic Chemistry Some Basic Principles and Techniques Unit XIII Hydrocarbons Classification of Hydrocarbons Unit XIV Environmental Chemistry Content 1 Some Basic Concepts of Chemistry 2 Structure of Atom 3 Classification of Elements and Periodicity in Properties 4 Chemical Bonding and Molecular Structure 5 States of Matter 6 Thermodynamics 7 Equilibrium 8 Redox Reactions 9 Hydrogen 10 s Block Elements 11 p Block Elements 12 Organic Chemistry Some Basic Principles and Techniques 13 Hydrocarbons 14 Environmental Chemistry I Appendix II Log antilog Table **Chemistry<sup>3</sup>** Andrew Burrows, Andrew Parsons, Gwen Pilling, Gareth Price, 2013-03-21 New to this Edition **Chemistry** Martin Stuart Silberberg, 1996 This general chemistry text centres on the theme that observable change in chemical systems is the result of molecular change The aims of this edition are to enable students to perceive matter and change at the molecular level and to help build student confidence in their ability to solve chemical problems as they discover the relevance of chemistry to their lives **Student Solutions Manual: Ssm Chemistry** Deborah Wiegand, 2003 This manual contains complete worked out solutions to all follow up problems and about half of all the chapter problems Each chapter of solutions opens with a summary of the text chapter content and a list of key equations needed to solve the problems **NCERT Chemistry Class 11 - [CBSE Board]** Dr. S. C. Rastogi, ,Er. Meera Goyal, 2022-10-11 Syllabus Unit I Some Basic Concepts of Chemistry Unit II Structure of Atom Unit III Classification of Elements and Periodicity in Properties Unit IV Chemical Bonding and Molecular

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**Quantitative Chemical Analysis** Daniel C. Harris, Chuck Lucy, 2015-05-29 The gold standard in analytical chemistry Dan Harris Quantitative Chemical Analysis provides a sound physical understanding of the principles of analytical chemistry and their applications in the disciplines *Student Solutions Manual* Deborah Wiegand, 1999-07

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