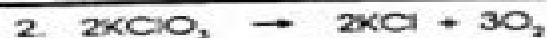


**STOICHIOMETRY:
MOLE-MOLE PROBLEMS**

How many moles of hydrogen are needed to completely react with two moles of nitrogen?

$$2 \text{ mol N}_2 \times \frac{3 \text{ mol H}_2}{1 \text{ mol N}_2} = 6$$

6 moles H₂

How many moles of oxygen are produced by the decomposition of six moles of potassium chlorate?

$$6 \text{ moles KClO}_3 \times \frac{3 \text{ mol O}_2}{2 \text{ mol KClO}_3} = \frac{18}{2} = 9$$

9 moles of O₂

How many moles of hydrogen are produced from the reaction of three moles of zinc with an excess of hydrochloric acid?

$$3 \text{ mol Zn} \times \frac{1 \text{ mol H}_2}{1 \text{ mol Zn}} = 3$$

3 moles of H₂How many moles of oxygen are necessary to react completely with four moles of propane (C₃H₈)?

$$4 \text{ mol C}_3\text{H}_8 \times \frac{5 \text{ mol O}_2}{1 \text{ mol C}_3\text{H}_8} = 20$$

20 mol of O₂

How many moles of potassium nitrate are produced when two moles of potassium phosphate react with two moles of aluminum nitrate?

$$2 \text{ mol K}_3\text{PO}_4 \times \frac{3 \text{ mol KNO}_3}{1 \text{ mol K}_3\text{PO}_4} = 6$$

6 moles of KNO₃

Moles And Stoichiometry Packet Answers

Sara McCubbins, Angela Codron



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Improving Student Comprehension of Stoichiometric Concepts Connie Lynn Bannick Kemner, 2007 **Forensics in Chemistry** Sara McCubbins, Angela Codron, 2012 Forensics seems to have the unique ability to maintain student interest and promote content learning I still have students approach me from past years and ask about the forensics case and specific characters from the story I have never had a student come back to me and comment on that unit with the multiple choice test at the end from the Introduction to Forensics in Chemistry The Murder of Kirsten K How did Kirsten K's body wind up at the bottom of a lake and what do wedding cake ingredients soil samples radioactive decay bone age blood stains bullet matching and drug lab evidence reveal about whodunit These mysteries are at the core of this teacher resource book which meets the unique needs of high school chemistry classes in a highly memorable way The book makes forensic evidence the foundation of a series of eight hands on week long labs As you weave the labs throughout the year and students solve the case the narrative provides vivid lessons in why chemistry concepts are relevant and how they connect All chapters include case information specific to each performance assessment and highlight the related national standards and chemistry content Chapters provide Teacher guides to help you set up Student performance assessments A suspect file to introduce the characters and new information about their relationships to the case Samples of student work that has been previously assessed and that serves as an answer key for you Grading rubrics Using Forensics in Chemistry as your guide you will gain the confidence to use inquiry based strategies and performance based assessments with a complex chemistry curriculum Your students may gain an interest in chemistry that rivals their fascination with Bones and CSI *Solving Problems in Chemistry* Rod O'Connor, Charles Mickey, Alton Hassell, 1977 *You Can Do Chemistry* Sunil Tanna, 2018-11-19 A comprehensive guide to performing mole and stoichiometric calculations with numerous examples as well as questions and answers Covers calculations relating to solids solutions gases and electrolysis plus as limiting and excess reactants chemical yields atom economy and much more Fully up to date with the last international standards including the revised definition of mole which was agreed on November 16th 2018 *Mole Concepts and Stoichiometry* Dharshi Bopegedera, 2017

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5 Steps Chemistry Workbook Series Book 2: Mighty Mole Concepts Julie C. Gilbert, 2022-03-07. Chemistry moles got you down? Mole concepts is a challenging unit because there are a lot of different topics. Whether you're a teacher looking for easy worksheets to borrow or a student wanting more practice, I've got something for you. Inside, you'll find:

- Descriptions for each of the major mole concepts
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- 4 worksheets covering various mole conversion topics
- 3 worksheets covering percent calculations
- 4 worksheets covering empirical, molecular, and moles of hydrates calculations
- 2 mixed moles self-tests with answer keys

This is a companion workbook for the 5 Steps to Surviving Chemistry book and the 5 Steps Chemistry Workbook Series Book 1: Stoichiometry. However, you do not need to have read those books to find this workbook useful.

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the key topics in basic stoichiometry including the mole concept reacting quantities and empirical and molecular formulas It begins with the basic concepts and formulas required to convert various quantities to moles or amount of substance This is particularly useful to a beginner Chapter 2 describes how to calculate reacting quantities and therefore provides a general step by step framework or approach by which to solve these problems The chapter also describes and applies the concept of a limiting reagent Two methods of determining a limiting reagent are explained and illustrated The concept of limiting reagent is extended to reacting volumes of gases with a short cut method A short cut method for solving reacting quantities involving masses and volumes of gases is also given Chapter 3 describes the calculations involved in the practical determination of molecular and empirical formulas A clear meaning of percentage composition of mass is provided and used to solve problems in a step by step manner In chapter 4 we discuss percentage yield and purity A number of examples are given to illustrate how formulas of yield and purity are used in various circumstances A student will find these examples helpful in relating different formulas for percentage purity The last chapter introduces a graphical method for reacting quantities The method may provide a new way of looking at chemical reactions Examples are given to illustrate the method including how it can be used to determine limiting reagents It is hoped that the book will provide all the necessary knowledge and skills to students studying an introductory chemistry course Teachers may also find this book a good resource for their lessons in stoichiometry

Teaching the Mole Concept in Stoichiometry Timothy Scot Albrecht, California State Polytechnic University, Pomona. College of Education & Integrative Studies, 2009 Math review, moles, stoichiometry, nomenclature, solutions, enthalpy, gas laws, atomic structure, Lewis structures, oxidation-reduction, colligative properties Drago. Russell S., 1984 **Chemistry Equations & Answers (Speedy Study Guide)** Speedy Publishing, 2014-11-26

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