

Limiting Reactant

For the following reactions, find the following:

- Which of the reactants is the limiting reactant?
 - What is the maximum amount of each product that can be formed?
 - How much of the excess reactant is left over after the reaction is complete?
- See next page for calculations!

1. Consider the following reaction:



a) limiting reactant

Answer the questions above, assuming we started with 30.0 grams of ammonium nitrate and 50.0 grams of sodium phosphate.

$$\begin{aligned} & \frac{30.0 \text{ g NH}_4\text{NO}_3}{1} \times \frac{1 \text{ mol NH}_4\text{NO}_3}{80.0 \text{ g NH}_4\text{NO}_3} \times \frac{1 \text{ mol } (\text{NH}_4)_3\text{PO}_4}{3 \text{ mol NH}_4\text{NO}_3} \times \frac{149.12 \text{ g } (\text{NH}_4)_3\text{PO}_4}{1 \text{ mol } (\text{NH}_4)_3\text{PO}_4} = 18.6 \text{ g } (\text{NH}_4)_3\text{PO}_4 \\ & \frac{50.0 \text{ g Na}_3\text{PO}_4}{1} \times \frac{1 \text{ mol Na}_3\text{PO}_4}{162.94 \text{ g Na}_3\text{PO}_4} \times \frac{1 \text{ mol } (\text{NH}_4)_3\text{PO}_4}{1 \text{ mol Na}_3\text{PO}_4} \times \frac{149.12 \text{ g } (\text{NH}_4)_3\text{PO}_4}{1 \text{ mol } (\text{NH}_4)_3\text{PO}_4} = 45.5 \text{ g } (\text{NH}_4)_3\text{PO}_4 \\ & \frac{30.0 \text{ g NH}_4\text{NO}_3}{1} \times \frac{1 \text{ mol NH}_4\text{NO}_3}{80.0 \text{ g NH}_4\text{NO}_3} \times \frac{3 \text{ mol NaNO}_3}{3 \text{ mol NH}_4\text{NO}_3} \times \frac{85.00 \text{ g NaNO}_3}{1 \text{ mol NaNO}_3} = 31.9 \text{ g NaNO}_3 \end{aligned}$$

2. Consider the following reaction:



a) limiting reactant

Answer the questions at the top of this sheet, assuming we start with 100.0 grams of calcium carbonate and 45.0 grams of iron (II) phosphate.

$$\begin{aligned} & \frac{100.0 \text{ g CaCO}_3}{1} \times \frac{1 \text{ mol CaCO}_3}{100.09 \text{ g CaCO}_3} \times \frac{1 \text{ mol Ca}_3(\text{PO}_4)_2}{3 \text{ mol CaCO}_3} \times \frac{310.18 \text{ g Ca}_3(\text{PO}_4)_2}{1 \text{ mol Ca}_3(\text{PO}_4)_2} = 103.3 \text{ g Ca}_3(\text{PO}_4)_2 \\ & \frac{45.0 \text{ g FePO}_4}{1} \times \frac{1 \text{ mol FePO}_4}{150.82 \text{ g FePO}_4} \times \frac{1 \text{ mol Ca}_3(\text{PO}_4)_2}{2 \text{ mol FePO}_4} \times \frac{310.18 \text{ g Ca}_3(\text{PO}_4)_2}{1 \text{ mol Ca}_3(\text{PO}_4)_2} = 46.3 \text{ g Ca}_3(\text{PO}_4)_2 \\ & \frac{45.0 \text{ g FePO}_4}{1} \times \frac{1 \text{ mol FePO}_4}{150.82 \text{ g FePO}_4} \times \frac{1 \text{ mol Fe}_2(\text{CO}_3)_3}{2 \text{ mol FePO}_4} \times \frac{291.73 \text{ g Fe}_2(\text{CO}_3)_3}{1 \text{ mol Fe}_2(\text{CO}_3)_3} = 43.5 \text{ g Fe}_2(\text{CO}_3)_3 \end{aligned}$$

Limiting Reactant Review Sheet Key

Martin Silberberg, Libby Weberg



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The Practice of Chemistry Study Guide & Solutions Manual Pamela Mills, Amina El-Ashmawy, 2003-04-14 Designed to help students understand the material better and avoid common mistakes Also includes solutions and explanations to odd numbered exercises

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CHEMISTRY HANDBOOK & STUDY GUIDE Gr11-12 NE Kevin Smith, 2024-02-01 A comprehensive summary of Grade 11 12 Physics Simple logical summaries with example exam questions and work through solutions The book covers the fundamentals of Grade 11 12 Physics and complements the material in any class text

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Study Guide for Introductory Chemistry Iris Stovall, 1990

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Foundations of College Chemistry, Alternate Morris Hein, Susan Arena, 2010-01-26 Learning the fundamentals of chemistry can be a difficult task to undertake for health professionals For over 35 years this book has helped them master the chemistry skills they need to succeed It provides them with clear and logical explanations of chemical concepts and problem solving They'll learn how to apply concepts with the help of worked out examples In addition Chemistry in Action features and conceptual questions checks brings together the understanding of chemistry and relates chemistry to things health professionals experience on a regular basis

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General Chemistry Key Concise Review Guide Sixth Edition Darrell D. Ebbing, 1999

Holt Chemistry Ralph Thomas Myers, 2004

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