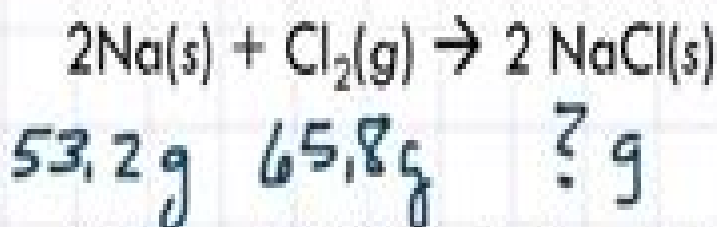


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- If we have 53.2 g of Na and 65.8 g of Cl_2 , what are the limiting reactant and theoretical yield?

$$53.2 \text{ g Na} \times \frac{1 \text{ mol Na}}{22.99 \text{ g Na}} \times \frac{2 \text{ mol NaCl}}{2 \text{ mol Na}} \times \frac{58.44 \text{ g NaCl}}{1 \text{ mol NaCl}} = 135.23$$

$$65.8 \text{ g Cl}_2 \times \frac{1 \text{ mol Cl}_2}{70.90 \text{ g Cl}_2} \times \frac{2 \text{ mol NaCl}}{1 \text{ mol Cl}_2} \times \frac{58.44 \text{ g NaCl}}{1 \text{ mol NaCl}} = 110.5$$

Limiting Reagent And Percent Yield Section Review

Morris Hein, Susan Arena



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compounds with structures that follow from simple methane found in natural gas to the complex macromolecules that serve as the basis of life on our planet The study of chemistry also follows from the simple to the more complex and the strength of this text is that it enables students with varied backgrounds to proceed together to significant levels of achievement

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Brady,John R. Holum,1993-02-11 Offers accurate lucid and interesting explanations of basic concepts and facts of chemistry while helping students develop skills in analytical thinking and problem solving Students are taught in a variety of ways to think of skills as tools that can be used to solve complex problems Several aids are included to help focus and inspire student interest frequent reference to common chemicals in commercial products numerous photographs of reactions in chapter practice exercises following worked examples **General Chemistry** James E. Brady,1986 *Ebook: Chemistry: The Molecular Nature of Matter and Change* Silberberg,2015-01-16 Ebook Chemistry The Molecular Nature of Matter and Change **Study Guide for Introductory Chemistry** Iris Stovall,1990 **Introductory Chemistry** Nivaldo J. Tro,2006

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