Limiting Reactant Worksheet #2

1. Given the following reaction: (Balance the equation first!)

 $C_3H_8 + O_2 - CO_2 + H_2O$

- a) If you start with 14.8 g of C₃H₈ and 3.44 g of O₂, determine the limiting reactant.
- b) Determine the number of moles of carbon dioxide produced.
- c) Determine the number of grams of H₂O produced.
- d) Determine the number of grams of excess reactant left.
- Given the following equation:

Al2(SO3)3 + 6NaOH -----> 3Na2SO3 + 2Al(OH)3

- a) If 10.0 g of Al₂(SO₂)₃ is reacted with 10.0 g of NaOH, determine the limiting reactant.
- b) Determine the number of moles of Al(OH)₃ produced.
- c) Determine the number of grams of Na₂SO₃ produced.
- d) Determine the number of grams of excess reactant left over in the reaction.
- 3. Given the following equation:

 $Al_2O_3 + Fe ----> Fe_3O_4 + Al$

- a) If 25.4 g of Al₂O₃ is reacted with 10.2 g of Fe, determine the limiting reactant.
- b) Determine the number of moles of Al produced.
- c) Determine the number of grams of Fe₃O₄ produced.
- d) Determine the number of grams of excess reactant left over in the reaction.
- Consider the reaction

 $I_2O_5(g) + 5 CO(g) -----> 5 CO_2(g) + I_2(g)$

- a) 80.0 grams of I₂O₅, reacts with 28.0 grams of carbon monoxide, CO. Determine the mass of iodine I₂, which could be produced?
- b) If, in the above situation, only 0.160 moles, of iodine, I₂ was produced.
 - i) What mass of iodine was produced?
 - ii) What percentage yield of iodine was produced?
- 5. Zinc and sulfur react to form zinc sulfide according to the equation

Zn + S -----> ZnS

If 25.0 g of zinc and 30.0 g of sulfur are mixed,

- a) Which chemical is the limiting reactant?
- b) How many grams of ZnS will be formed?
- c) How many grams of the excess reactant will remain after the reaction is over?
- 6. Which element is in excess when 3.00 grams of Mg is ignited in 2.20 grams of pure oxygen? What mass is in excess? What mass of MgO is formed?
- 7. How many grams of Al₂S₃ are formed when 5.00 grams of Al is heated with 10.0 grams S?

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