

AP Chemistry Chapter 14 Answers – Zumdahl

14.27

a.
 $\text{HClO}_4(\text{aq}) + \text{H}_2\text{O}(\text{l}) \rightarrow \text{H}_3\text{O}^+(\text{aq}) + \text{ClO}_4^-(\text{aq})$. Only the forward reaction is indicated since HClO_4 is a strong acid and is basically 100% dissociated in water. For acids, the dissociation reaction is commonly written without water as a reactant. The common abbreviation for this reaction is: $\text{HClO}_4(\text{aq}) \rightarrow \text{H}^+(\text{aq}) + \text{ClO}_4^-(\text{aq})$. This reaction is also called the K_a reaction as the equilibrium constant for the reaction is called K_a .

b.
 Propanoic acid is a weak acid, so it is only partially dissociated in water. The dissociation reaction is: $\text{CH}_3\text{CH}_2\text{CO}_2\text{H}(\text{aq}) + \text{H}_2\text{O}(\text{l}) \rightleftharpoons \text{H}_3\text{O}^+ + \text{CH}_3\text{CH}_2\text{CO}_2^-(\text{aq})$ or $\text{CH}_3\text{CH}_2\text{CO}_2\text{H}(\text{aq}) \rightleftharpoons \text{H}^+(\text{aq}) + \text{CH}_3\text{CH}_2\text{CO}_2^-(\text{aq})$.

c.
 NH_4^+ is a weak acid. Similar to propanoic acid, the dissociation reaction is:



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An acid is a proton (H^+) donor and a base is a proton acceptor. A conjugate acid-base pair differs by only a proton (H^+).

	Acid	Base	Conjugate Base of Acid	Conjugate Acid of Base
a.	H_2CO_3	H_2O	HCO_3^-	H_3O^+
b.	$\text{C}_5\text{H}_5\text{NH}^+$	H_2O	$\text{C}_5\text{H}_5\text{N}$	H_3O^+
c.	$\text{C}_4\text{H}_9\text{NH}^+$	HCO_3^-	$\text{C}_4\text{H}_9\text{N}$	H_2CO_3

14.31

Strong acids have a $K_a \gg 1$ and weak acids have $K_a < 1$. Table 14.2 in the text lists some K_a values for weak acids. K_a values for strong acids are hard to determine so they are not listed in the text. However, there are only a few common strong acids so if you memorize the strong acids, then all other acids will be weak acids. The strong acids to memorize are HCl , HBr , HI , HNO_3 , HClO_4 , and H_2SO_4 .

- HClO_4 is a strong acid.
- HOCl is a weak acid ($K_a = 3.5 \times 10^{-8}$).
- H_2SO_4 is a strong acid.
- H_2SO_3 is a weak diprotic acid with K_{a1} and K_{a2} values less than one.

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The K_a value is directly related to acid strength. As K_a increases, acid strength increases. For water, use K_w when comparing the acid strength of water to other species. The K_a values are:



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