

Half Modern Chemistry Review

CHAPTER 9: STOICHIOMETRY

The following pages contain the book (but not all) of the information for the chapter 9 test.

Focus on this content, but make sure to review class notes, activities, assessments, questions, etc.

If you study this document and NOTHING else, you should at least be able to PASS the test.

*** Test items will be recall, examples, and/or application of this content. ***

OUTCOMES

- Calculate molar percent or molar-mass-molar percent (M/M)
- Calculate the molar-mass-molar percent (M/M) and percent (M/M)
- 9.1. Determine percent of water from balanced chemical equations. (T & P)
- 9.2. Perform stoichiometry calculations with: a) mole to mole, mole to gram, gram to mole, and gram to gram. (P & P)
- 9.3. Calculate stoichiometry problems and determine the limiting reagent, excess reagent, and amount of product produced. (P & P)

9.1: INTRODUCTION TO STOICHIOMETRY

a. Introductory

- stoichiometry calculations -- calculations involving the mass relationships of elements in compounds.
- molar-mass-molar calculations -- calculations involving the mass relationships between reactants and products in a chemical reaction.
- molar ratio -- a conversion factor that relates the amount of any one substance involved in a chemical reaction.

b. Chapter Highlights

- Stoichiometry calculations involve the mass relationships between reactants and products in a chemical reaction.
- Stoichiometry is based on the law of conservation of mass.
- Stoichiometry is based on the molar ratio and molar mass of a substance in a reaction.
- A molar ratio is the conversion factor that relates the amount in moles of any two substances in a chemical reaction. The molar ratio is derived from the balanced equation.
 - Example: $2\text{FeCl}_3 + 3\text{H}_2\text{O} \rightarrow \text{Fe}_2\text{O}_3 + 6\text{HCl}$ (Note: the molar ratio of FeCl_3 to Fe_2O_3 is 2 to 1)
- Amount of a substance is measured in moles, and mass of a substance is determined by using molar mass, such as grams, kilograms, or milligrams.
- Mass and amount of substances are quantities, substances are also quantities, and mass and amount of any other quantity is derived from the balanced equation.

9.2: MASS STOICHIOMETRIC CALCULATIONS

a. Chapter Highlights

- In an ideal stoichiometric calculation, the mass of the amount of any reactant or product can be calculated if the balanced chemical equation and the molar mass of any other quantity is known.

9.3: LIMITING REACTANTS AND PERCENTAGE YIELD

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