

balance nuclear reaction equations

aluminum-26 undergoes positron emission
potassium-40 undergoes beta decay (electron)
uranium-238 undergoes alpha decay



electron configurations

Write the electron configuration for C.

Write the electron configuration for Cr.

Write the shorthand notation for Bi.



NOMENCLATURE [naming compounds, writing formulas]

Write the formula for each of the following compounds:

- a. Sulfur hexafluoride SF_6
- b. Lithium nitride Li_3N
- c. Chromium (III) carbonate $\text{Cr}_2(\text{CO}_3)_3$
- d. Tin (II) chloride SnCl_2
- e. Ammonium acetate $\text{NH}_4\text{C}_2\text{H}_3\text{O}_2$ (or $\text{CH}_3\text{COONH}_4$)
- f. Mercury (I) chloride Hg_2Cl_2
- g. Potassium bromate KBrO_3
- h. Hydrosulfuric acid $\text{H}_2\text{S(aq)}$
- i. Chloric acid HClO_3
- j. Sulfurous acid H_2SO_3

Name each of the following compounds.

- a. CuSO_4 Copper (II) sulfate
- b. AlF_3 aluminum fluoride
- c. HI hydrogen iodide
- d. NO nitrogen monoxide
- e. $\text{H}_2\text{Se(aq)}$ hydroselenic acid
- f. HNO_3 nitric acid
- g. HNO_2 nitrous acid
- h. NaHSO_4 sodium hydrogen sulfate

Half-life

Tritium (H-3) is a radioactive isotope of hydrogen with a half-life of 12.3 years. How long would it take for a 40.0 g sample to decay down to 1.25 g?

$$40 \xrightarrow{1} 20 \xrightarrow{2} 10 \xrightarrow{3} 5 \xrightarrow{4} 2.5 \xrightarrow{5} 1.25$$

Fe-61 has a half-life of 6.00 min. Of a 100.0 mg sample, how much will remain after 18.0 min?

$$100 \xrightarrow{1} 50 \xrightarrow{2} 25 \xrightarrow{3} 12.5 \text{ mg}$$

3 half lives

Average atomic mass

The element copper has naturally occurring isotopes with mass numbers of 63 and 65.

The relative abundance and atomic masses are 69.2% for a mass of 63amu and 30.8% for a mass of 65amu. Calculate the average atomic mass of copper.

$$\begin{array}{rcl} 63 \times .692 & = & 43.596 \rightarrow 44 \\ 65 \times .308 & = & 20.02 \rightarrow 20 \\ \hline & + & \\ & 63.62 & \rightarrow 64 \end{array}$$

Modern Chemistry Mixed Review Chapter7 Answers

N Colangelo



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Chemistry Harold Clark Metcalfe, 1982 **Modern Electrochemistry 2B** John O'M. Bockris, Amulya K.N. Reddy, 2007-05-08 This book had its nucleus in some lectures given by one of us J O M B in a course on electrochemistry to students of energy conversion at the University of Pennsylvania It was there that he met a number of people trained in chemistry physics biology metallurgy and materials science all of whom wanted to know something about electrochemistry The concept of writing a book about electrochemistry which could be understood by people with very varied backgrounds was thereby engendered The lectures were recorded and written up by Dr Klaus Muller as a 293 page manuscript At a later stage A K N R joined the effort it was decided to make a fresh start and to write a much more comprehensive text Of methods for direct energy conversion the electrochemical one is the most advanced and seems the most likely to become of considerable practical importance Thus conversion to electrochemically powered transportation systems appears to be an important step by means of which the difficulties of air pollution and the effects of an increasing concentration in the atmosphere of carbon dioxide may be met Corrosion is recognized as having an electrochemical basis The synthesis of nylon now contains an important electrochemical stage Some central biological mechanisms have been shown to take place by means of electrochemical reactions A number of American organizations have recently recommended greatly increased activity in training and research in electrochemistry at universities in the United States *Applied Polymer Science* Ulf W. Gedde, Mikael S. Hedenqvist, Minna Hakkarainen, Fritjof Nilsson, Oisik Das, 2021-10-29 This companion volume to *Fundamental Polymer Science* Gedde and Hedenqvist 2019 offers detailed insights from leading practitioners into experimental methods simulation and modelling mechanical and transport properties processing and sustainability issues

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Colliery Guardian, and Journal of the Coal and Iron Trades ,1922 **Scientific American** ,1894 *The Chemical Age* ,1928 *Mining and Engineering World* ,1907 *Gardeners' Chronicle* ,1857 **New Technical Books** New York Public Library,1993 *The Gardeners' Chronicle and Agricultural Gazette* ,1857 *Progress in Modern Hydrology* John C. Rodda,Mark Robinson,2015-08-26 Hydrology is vital to human civilisations as well as to natural ecosystems yet it has only emerged as a distinct scientific discipline during the last 50 years or so This book reviews the development of modern hydrology primarily through the experiences of the multidisciplinary team of scientists and engineers at Wallingford near Oxford who have been at the forefront of many of the developments in UK hydrological research These topics include The development of basic understanding through the collection of data with specialised instrumentation in experimental basins The study of extreme flows both floods and droughts The role moisture in the soil Studies of the processes controlling evaporation Water resource studies Modelling and prediction of the extremes of flow improved Understanding of water quality issues A widening recognition of the importance of an ecosystem approach Meeting the challenges of climate change Data handling Future developments in hydrology and the pressures which generate them Readership hydrologists in both academia and a wide range of applied fields such as civil engineering meteorology geography and physics as well as advanced students in earth science environmental science and physical geography programmes worldwide Technical Book Review Index ,1973

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