

## Practice Test Ch 3

## Stoichiometry

Name: \_\_\_\_\_ Per: \_\_\_\_\_

- Remember, this is practice - Do NOT cheat yourself of finding out what you are capable of doing. Be sure you follow the testing conditions outlined below.
- DO NOT USE A CALCULATOR. AP Chem does not allow the use of a calculator for the MC part of the exam, so it is time to start practicing without one. You may use ONLY a periodic table.
- While you should practice working as fast as possible, it is more important at this point in the course, that you practice without a calculator, even if it slows you down. Look for the "easy math" – common factors and rough estimation – do not do "long division" to try to get exact values. Remember it is a MC test, use the answers
- Mark which questions you would like to "go over" when we get to school in September.

1. Balance the following equation:



The balanced equation shows that 1.00 mole of  $\text{NH}_3$  requires \_\_\_\_\_ mole(s) of  $\text{O}_2$

- 0.57
- 1.25
- 1.33
- 1.75
- 3.5

2. Write a balanced equation for the combustion of acetaldehyde,
- $\text{CH}_3\text{CHO}$
- .

When properly balanced, the equation indicates that \_\_\_\_\_ mole(s) of  $\text{O}_2$  are required for each mole of  $\text{CH}_3\text{CHO}$ .

- 1
- 2
- 2.5
- 3
- 5

3. What is the total mass of products formed when 16 grams of
- $\text{CH}_4$
- is burned with excess oxygen?

- 32 g
- 36 g
- 44 g
- 62 g
- 80 g

4. Write a balanced equation for the combustion of propane,
- $\text{C}_3\text{H}_8$
- .

When balanced, the equation indicates that \_\_\_\_\_ moles of  $\text{O}_2$  are required for each mole of  $\text{C}_3\text{H}_8$ .

- 1.5
- 3
- 3.5
- 5
- 8

5. Balance the following equation with the SMALLEST WHOLE NUMBER COEFFICIENTS possible. Select the number that is the
- sum
- of the coefficients in the balanced equation:



- 3
- 5
- 6
- 7
- 8

6. Calculate the mass of hydrogen formed when 27 g of aluminum reacts with excess hydrochloric acid according to the balanced equation below.



- 1.5 g
- 2.0 g
- 3.0 g
- 6.0 g
- 12 g

7. How many grams of nitric acid,
- $\text{HNO}_3$
- , can be prepared from the reaction of 138 g of
- $\text{NO}_2$
- with 54.0 g
- $\text{H}_2\text{O}$
- according to the equation below?



- 92
- 108
- 126
- 189
- 279

8. Which of the following statements is true?

- The molar mass of  $\text{CaCO}_3$  is  $100.1 \text{ g mol}^{-1}$ .
- 50 g of  $\text{CaCO}_3$  contains  $9 \times 10^{23}$  oxygen atoms.
- A 200 g sample of  $\text{CaCO}_3$  contains 2 moles of  $\text{CaCO}_3$ .

- I only
- II only
- III only
- I and III only
- I, II, and III

# Modern Chemistry Stoichiometry Test Answers

**Ian Pickup**



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