

Experiment 7



Molar Mass Determination

An Introduction to Gas Laws

In this experiment, the molar mass of butane is determined using the ideal gas law and Dalton's law of partial pressures.

Learning Objectives: Students will

- Learn how to use the ideal gas law and Dalton's law of partial pressures to calculate the molar mass of a gas
- Practice collection of a gas using water displacement

Experiment Objectives:

- Calculate the molar mass of butane
- Statistically determine the accuracy and precision of the results

Introduction

Context: Texas Natural Gas

In 2009, Texas consumed 3.384 billion cubic feet of natural gas, more than any other state in the nation.¹ Natural gas is a fossil fuel composed primarily of methane, although significant amounts of ethane, propane, and butane are present. Carbon dioxide, helium, hydrogen sulfide, and nitrogen gases are found in lesser amounts. During processing, the various gases are separated to form wet natural gas and dry natural gas. Wet natural gas is composed of hydrocarbons other than methane. Dry natural gas is composed of only methane. Household natural gas is dry natural gas. Wet natural gas is further refined, yielding ethane, propane, and butane gases. Butane gas is the subject of this lab.

Component	Typical wt. %
Methane (CH_4)	70–90
Ethane (C_2H_6)	5–15
Propane (C_3H_8) and Butane (C_4H_{10})	<5
CO_2 , N_2 , H_2S , etc.	balance

Molar Mass Determination Real World Applications

Sabine Zange

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