

Dilutions Worksheet - Solutions

- 1) If I add 25 mL of water to 125 mL of a 0.15 M NaOH solution, what will the molarity of the diluted solution be?

$$\begin{aligned}M_1V_1 &= M_2V_2 \\(0.15\text{ M})(125\text{ mL}) &= x(150\text{ mL}) \\x &= 0.125\text{ M}\end{aligned}$$

- 2) If I add water to 100 mL of a 0.15 M NaOH solution until the final volume is 150 mL, what will the molarity of the diluted solution be?

$$\begin{aligned}M_1V_1 &= M_2V_2 \\(0.15\text{ M})(100\text{ mL}) &= x(150\text{ mL}) \\x &= 0.100\text{ M}\end{aligned}$$

- 3) How much 0.05 M HCl solution can be made by diluting 250 mL of 10 M HCl?

$$\begin{aligned}M_1V_1 &= M_2V_2 \\(10\text{ M})(250\text{ mL}) &= (0.05\text{ M})x \\x &= 50,000\text{ mL}\end{aligned}$$

- 4) I have 345 mL of a 1.5 M NaCl solution. If I boil the water until the volume of the solution is 250 mL, what will the molarity of the solution be?

$$\begin{aligned}M_1V_1 &= M_2V_2 \\(1.5\text{ M})(345\text{ mL}) &= x(250\text{ mL}) \\x &= 2.07\text{ M}\end{aligned}$$

- 5) How much water would I need to add to 500 mL of a 2.4 M KCl solution to make a 1.0 M solution?

$$\begin{aligned}M_1V_1 &= M_2V_2 \\(2.4\text{ M})(500\text{ mL}) &= (1.0\text{ M})x \\x &= 1200\text{ mL}\end{aligned}$$

1200 mL will be the final volume of the solution. However, since there's already 500 mL of solution present, you only need to add 700 mL of water to get 1200 mL as your final volume. The answer: 700 mL.

Molarity By Dilution Page 69 Answers

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Table of Contents Molarity By Dilution Page 69 Answers

1. Understanding the eBook Molarity By Dilution Page 69 Answers
 - The Rise of Digital Reading Molarity By Dilution Page 69 Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Molarity By Dilution Page 69 Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Molarity By Dilution Page 69 Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Molarity By Dilution Page 69 Answers
 - Personalized Recommendations
 - Molarity By Dilution Page 69 Answers User Reviews and Ratings
 - Molarity By Dilution Page 69 Answers and Bestseller Lists
5. Accessing Molarity By Dilution Page 69 Answers Free and Paid eBooks

- Molarity By Dilution Page 69 Answers Public Domain eBooks
- Molarity By Dilution Page 69 Answers eBook Subscription Services
- Molarity By Dilution Page 69 Answers Budget-Friendly Options
- 6. Navigating Molarity By Dilution Page 69 Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Molarity By Dilution Page 69 Answers Compatibility with Devices
 - Molarity By Dilution Page 69 Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Molarity By Dilution Page 69 Answers
 - Highlighting and Note-Taking Molarity By Dilution Page 69 Answers
 - Interactive Elements Molarity By Dilution Page 69 Answers
- 8. Staying Engaged with Molarity By Dilution Page 69 Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Molarity By Dilution Page 69 Answers
- 9. Balancing eBooks and Physical Books Molarity By Dilution Page 69 Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Molarity By Dilution Page 69 Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Molarity By Dilution Page 69 Answers
 - Setting Reading Goals Molarity By Dilution Page 69 Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Molarity By Dilution Page 69 Answers
 - Fact-Checking eBook Content of Molarity By Dilution Page 69 Answers
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development

- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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die neuordnungsdebatte 61

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