

# MOLE CONCEPT

## CLASS X<sup>th</sup>

1. Calculate the relative molecular masses (or molecular weights) of the following compounds :

- i) Copper sulphate crystals,  $\text{CuSO}_4 \cdot 5\text{H}_2\text{O}$ .
- ii) Ammonium sulphate,  $(\text{NH}_4)_2\text{SO}_4$ .
- iii) Cane sugar,  $\text{C}_{12}\text{H}_{22}\text{O}_{11}$ .

(Given atomic masses Cu = 63.5, S = 32, O = 16, N = 14, C = 12)

2. An organic compound with vapour density = 94 contains. C = 12.67%, H = 2.13%, and Br = 85.11%. Find the molecular formula. [Atomic mass : C = 12, H = 1, Br = 80]

3. (i) The compound A has the following percentage composition by mass. Carbon = 26.7%, oxygen = 71.1%, hydrogen = 2.2%. Determine the Empirical formula of compound A [work to one decimal place] [H = 1, C = 12, O = 16].

ii) If the relative molecular mass of A is 90, what is the molecular formula of A?

(iii) The compound A is weak acid. What is meant by this statement?

4. Determine the empirical formula of the compound whose composition by mass is: 42% nitrogen, 48% oxygen and 9% hydrogen. [H = 1; N = 14; O = 16]

5. Determine the empirical formula of a compound containing 47.9% potassium, 5.5% beryllium and 46.6% fluorine by mass. (Atomic weight of Be = 9; F = 19; K = 39) Work to one decimal place.

6. The Empirical formula of a compound is  $\text{C}_2\text{H}_5$ . It has a vapour density of 29. Determine the relative molecular formula mass of the compound and hence its molecular formula.

7. The percentage composition of sodium phosphate as determined by analysis, is 42.1% sodium, 18.9% phosphorus and 39% oxygen. Find the empirical formula of the compound (work to two decimal places).

8. A gaseous organic compound contains 3.6 g of carbon and 0.8 g of hydrogen. The vapour density of this compound is 22. (i) Calculate the Empirical formula. (ii) Calculate the molecular formula of the compound.

# Mole Concept Chemistry Puzzles Answers

**Martha Orrico Visscher**



## Mole Concept Chemistry Puzzles Answers:

**Chemistry and Industry**, 2007      **Laboratory Manual to Accompany The Ideas of Chemistry** Martha Orrico Visscher, 1978      **Chemistry, Problem-Solving Worktext** James E. Brady, Frederick Senese, 2002-06-13 The images on the cover call attention to the relationship between macro observations and the intimate structure of chemical substances and the changes both chemical and physical that they undergo Fireworks One of the ingredients is phosphorus a molecular form of which is believed to consist of linked tetrahedra of phosphorus atoms The chemical reaction of phosphorus with oxygen is partly responsible for the spectacular show of light Carbon The element is found in several forms including the familiar diamond and another recently discovered sooty substance that consists of soccer ball shaped molecules often referred to as buckeyballs Diamond is not the most stable form of carbon and is created from other forms of carbon at high temperatures and pressures deep within the earth Acetylene torch Cutting steel is possible because of the intense heat generated by the chemical reaction of acetylene with oxygen a reaction between molecules of  $C_2H_2$  and  $O_2$  to give  $CO_2$  and  $H_2O$  Hot air balloon The air that helps it rise is heated by the combustion of molecules of propane each composed of three carbon and eight hydrogen atoms Stormy weather The evaporation of water serves to store energy provided by the sun Subsequent condensation of the water vapor releases this energy and is the basis of all the weather systems on our planet      The Extraordinary Chemistry of Ordinary Things Carl H. Snyder, 1995-02-07 Examines the chemistry of the substances of our everyday world Our daily lives are immersed in chemicals an effective way to teach and learn chemistry is by examining the goods and substances that we use in our daily lives and that affect us and our environment      Holt Chemistry Ralph Thomas Myers, 2004      **General Chemistry** Ralph H. Petrucci, 1989      **Chemistry & Chemical Reactivity** John C. Kotz, Paul Treichel, Patrick A. Harman, 2003 Provides a broad overview of the principles of chemistry the reactivity of chemical elements and their compounds and the applications of chemistry Conveys a sense of chemistry as a field that not only has a lively history but also one that is currently dynamic with important new developments on the horizon      **Chemistry** Raymond Chang, 1988      *Quill & Quire*, 1962      Keeping the Faith Thomas J. Schoenbaum, 2015-03-10 This book explores the modern relationship between religion and science highlighting the public backlash over teaching evolution in U S public schools as an example of the conflict between religious faith and scientific reason The author explores the middle ground between the two most extreme arguments arguing that it is possible for science and religion to exist in harmony The book also provides an analysis of 10 world religions including Christianity Islam Judaism Hinduism and Buddhism revealing how each finds compatibility with differing religions and scientific doctrine      **Biology/science Materials** Carolina Biological Supply Company, 1991      **Catalog of Copyright Entries. Third Series** Library of Congress. Copyright Office, 1965 Includes Part 1 Number 2 Books and Pamphlets Including Serials and Contributions to Periodicals July December      **Chemistry** John S. Phillips, Cheryl Wistrom, 2000      *Chemistry* John McMurry, 1998      The Software Encyclopedia

,1988     *Harper's Weekly* John Bonner,George William Curtis, Henry Mills Alden, Samuel Stillman Conant, Montgomery Schuyler, John Foord, Richard Harding Davis, Carl Schurz, Henry Loomis Nelson, John Kendrick Bangs, George Brinton McClellan Harvey, Norman Hapgood, 1913     *Scientific American* ,1870     **American Book Publishing Record** ,1979  
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