

## MOLE RATIOS IN CHEMICAL EQUATIONS

When the equation is balanced, you can use the **coefficients** to get the **mole ratios** between all substances present to get the **amount** (moles) of each (pure) substance.

Remember, that one mole of a substance is equal to Avogadro's number ( $N_A$ ) =  $6.022 \times 10^{23}$  particles/mol of **particles** or **entities** of that substance (which could be a reactant or product). Below is a list of what an 'entity' represents for each type of pure substance.

**Element** – atom e.g. Al

**Covalent/Molecular Compound** – molecule e.g.  $H_2SO_4$

**Ionic Compound** – formula unit e.g.  $Al_2(SO_4)_3$

There is more than one way to set up/look at a mole ratio, but ultimately the final answer should be the same. There are short cuts too, but you should only use these after you understand what is happening.



### MOLE RATIOS



so, if you have 5.00 moles of Al, then how many moles of  $H_2SO_4$  are required to react with the Al?

*"cross multiplication" method*

$$\begin{array}{l} 2 \text{ Al} : 3 \text{ H}_2\text{SO}_4 \\ \frac{2 \text{ Al}}{5.00 \text{ Al}} = \frac{3 \text{ H}_2\text{SO}_4}{x \text{ mol H}_2\text{SO}_4} \end{array}$$

$$= \frac{3 \text{ H}_2\text{SO}_4 \times 5.00 \text{ Al}}{2 \text{ Al}}$$

*G x R/G method*

$$\begin{array}{c} G \times \frac{R}{G} = 5.00 \text{ Al} \times \frac{3 \text{ H}_2\text{SO}_4}{2 \text{ Al}} \\ \uparrow \\ \text{(coefficients)} \end{array} = 7.50 \text{ mols H}_2\text{SO}_4$$

# Mole Ratios Model 1 A Chemical Reaction

**Steven S. Zumdahl**



## **Mole Ratios Model 1 A Chemical Reaction:**

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## **Table of Contents Mole Ratios Model 1 A Chemical Reaction**

1. Understanding the eBook Mole Ratios Model 1 A Chemical Reaction
  - The Rise of Digital Reading Mole Ratios Model 1 A Chemical Reaction
  - Advantages of eBooks Over Traditional Books
2. Identifying Mole Ratios Model 1 A Chemical Reaction
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Mole Ratios Model 1 A Chemical Reaction
  - User-Friendly Interface
4. Exploring eBook Recommendations from Mole Ratios Model 1 A Chemical Reaction
  - Personalized Recommendations
  - Mole Ratios Model 1 A Chemical Reaction User Reviews and Ratings

- Mole Ratios Model 1 A Chemical Reaction and Bestseller Lists
- 5. Accessing Mole Ratios Model 1 A Chemical Reaction Free and Paid eBooks
  - Mole Ratios Model 1 A Chemical Reaction Public Domain eBooks
  - Mole Ratios Model 1 A Chemical Reaction eBook Subscription Services
  - Mole Ratios Model 1 A Chemical Reaction Budget-Friendly Options
- 6. Navigating Mole Ratios Model 1 A Chemical Reaction eBook Formats
  - ePub, PDF, MOBI, and More
  - Mole Ratios Model 1 A Chemical Reaction Compatibility with Devices
  - Mole Ratios Model 1 A Chemical Reaction Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Mole Ratios Model 1 A Chemical Reaction
  - Highlighting and Note-Taking Mole Ratios Model 1 A Chemical Reaction
  - Interactive Elements Mole Ratios Model 1 A Chemical Reaction
- 8. Staying Engaged with Mole Ratios Model 1 A Chemical Reaction
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Mole Ratios Model 1 A Chemical Reaction
- 9. Balancing eBooks and Physical Books Mole Ratios Model 1 A Chemical Reaction
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Mole Ratios Model 1 A Chemical Reaction
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Mole Ratios Model 1 A Chemical Reaction
  - Setting Reading Goals Mole Ratios Model 1 A Chemical Reaction
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Mole Ratios Model 1 A Chemical Reaction
  - Fact-Checking eBook Content of Mole Ratios Model 1 A Chemical Reaction
  - Distinguishing Credible Sources

13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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