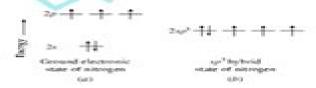
Oxygen is more electronegative than carbon and attracts electrons from it. Formaldehyde has a dipole moment.



(f) Nitrogen is more electronegative than carbon. Hydrogen cy. side has a dipole moment.



1.20 The orbital diagram for sp³-hybrid red is open is the same as for sp³-hybridized carbon, except nitrogen has one more electron.



The unshared electron pair in ammonia (NH₃) occupies an sp³-hybridized orbital of nitrogen. Each N—H bond corresponds to overlap of a half-filled sp³ hybrid orbital of nitrogen and a 1s orbital of hydrogen.

1.21 Silicon lies below carbon in the periodic table, and it is reasonable to assume that both carbon and silicon are sp²-hybridized in H₃CSiH₃. The C—Si bond and all of the C—H and Si—H bonds are σ bonds.

$$C(2sp^3) \longrightarrow B(1s) \Leftrightarrow bood \qquad H \longrightarrow H \qquad Sc(3sp^3) \longrightarrow B(1s) \Leftrightarrow bood$$

$$H \longrightarrow H \qquad H \qquad Sc(3sp^3) \longrightarrow B(1s) \Leftrightarrow bood$$

The principal quantum number of the carbon orbitals that are hybridized is 2; the principal quantum number for the silicon orbitals is 3.

- 1.22 (b) Carbon in formaldehyde (H₂C—O) is directly bonded to three other atoms (two hydrogens and one oxygen). It is sp²-hybridized.
 - (c) Ketene has two carbons in different hybridization states. One is xp³-hybridized; the other is xp-hybridized.



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