

Name _____ Date _____ Class _____

CHAPTER 12**STUDY GUIDE FOR CONTENT MASTERY****Section 12.3 Limiting Reactants**

In your textbook, read about why reactions stop and how to determine the limiting reactant.

Study the diagram showing a chemical reaction and the chemical equation that represents the reaction. Then complete the table. Show your calculations for questions 25–27 in the space below the table.



The molar masses of O_2 , NO , and NO_2 are 32.00 g/mol, 30.01 g/mol, and 46.01 g/mol, respectively.

Amount of O_2	Amount of NO	Amount of NO_2	Limiting Reactant	Amount and Mass of Excess Reactant
1 molecule	2 molecules	2 molecules	none	none
2 molecules	4 molecules	4 molecules	NO	2 molecules O_2
2 molecules	4 molecules	4 molecules	1. O_2	2. 4 molecules NO
1.00 mol	2.00 mol	2.00 mol	3. none	4. none
4.00 mol	6.00 mol	4.00 mol	5. NO	6. 2.00 mol O_2
6.00 mol	12.00 mol	6.00 mol	7. NO	8. 1.50 mol O_2
1.00 mol	6.00 mol	12.00 mol	9. O_2	10. 5.00 mol NO
6.000 mol	12.000 mol	12.000 mol	11. NO	12. 6.000 mol O_2
32.00 g	60.02 g	92.02 g	13. none	14. none
64.00 g	60.02 g	15. 46.01 g	15. O_2	16. 30.01 g NO
64.00 g	120.04 g	17. 28.76 g	17. O_2	18. 1.24 g NO

balanced equation mole ratio = 2 mol NO /1 mol O_2

$16.00 \text{ g } \text{O}_2 \times 1 \text{ mol } \text{O}_2 / 32.00 \text{ g } \text{O}_2 = 0.5125 \text{ mol } \text{O}_2$

$20.00 \text{ g } \text{NO} \times 1 \text{ mol } \text{NO} / 30.01 \text{ g } \text{NO} = 0.6664 \text{ mol } \text{NO}$

actual mole ratio = $0.6664 \text{ mol } \text{NO} / 0.5125 \text{ mol } \text{O}_2 = 1.302 \text{ mol } \text{NO} / 1.000 \text{ mol } \text{O}_2$

Because the actual mole ratio of $\text{NO}:\text{O}_2$ is larger than the balanced equation mole ratio of $\text{NO}:\text{O}_2$, there is an excess of NO ; O_2 is the limiting reactant.

Mass of NO used = $0.5125 \text{ mol } \text{O}_2 \times 2 \text{ mol } \text{NO} / 1 \text{ mol } \text{O}_2 = 0.4250 \text{ mol } \text{NO}$

$0.4250 \text{ mol } \text{NO} \times 30.01 \text{ g } \text{NO} / 1 \text{ mol } \text{NO} = 12.76 \text{ g } \text{NO}$

Mass of NO_2 produced = $0.4250 \text{ mol } \text{NO}_2 \times 46.01 \text{ g } \text{NO}_2 / 1 \text{ mol } \text{NO}_2 = 19.56 \text{ g } \text{NO}_2$

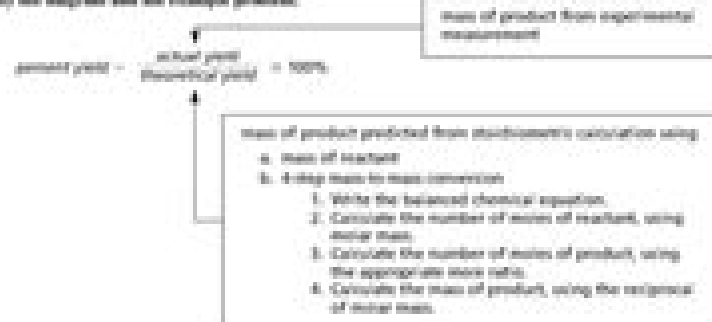
Excess $\text{NO} = 20.00 \text{ g } \text{NO} - 12.76 \text{ g } \text{NO} = 7.24 \text{ g } \text{NO}$

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CHAPTER 13**STUDY GUIDE FOR CONTENT MASTERY****Section 13.4 Percent Yield**

In your textbook, read about the yields of products.

Study the diagram and the example problem.



Example Problem: The following chemical equation represents the production of palladium, a valuable metal in the manufacturing of some semiconductor devices.



In one experiment, the reaction yielded 7.42 g of the metal from a 7.00-g sample of palladium. Determine the percent yield of this reaction. The molar masses of Ga and Ga_2O_3 are 69.72 g/mol and 157.44 g/mol, respectively.

Use the information in the diagram and example problem to evaluate each value or expression below. If the value or expression is correct, write correct. If it is incorrect, write the correct value or expression.

- actual yield unknown 7.42 g Ga_2O_3
- mass of reactant: 7.00 g Ga correct
- number of moles of reactant: $7.00 \text{ g } \text{Ga} \times \frac{69.72 \text{ g } \text{Ga}}{1 \text{ mol } \text{Ga}} = 1 \text{ mol } \text{Ga}$ 7.00 g $\text{Ga} \times 1 \text{ mol } \text{Ga} / 69.72 \text{ g } \text{Ga}$
- number of moles of product: $0.100 \text{ mol } \text{Ga} \times \frac{2 \text{ mol } (\text{Ga}_2\text{O}_3)}{1 \text{ mol } \text{Ga}} = 0.200 \text{ mol } \text{Ga}_2\text{O}_3$ 0.100 mol $\text{Ga} \times \frac{2 \text{ mol } \text{Ga}_2\text{O}_3}{1 \text{ mol } \text{Ga}}$
- theoretical yield: $0.100 \text{ mol } (\text{Ga}_2\text{O}_3) \times \frac{157.44 \text{ g } (\text{Ga}_2\text{O}_3)}{1 \text{ mol } (\text{Ga}_2\text{O}_3)} = 15.74 \text{ g } \text{Ga}_2\text{O}_3$ correct
- percent yield: $\frac{7.42 \text{ g } (\text{Ga}_2\text{O}_3)}{15.74 \text{ g } (\text{Ga}_2\text{O}_3)} \times 100 = 47.1\%$ $\frac{7.42 \text{ g } \text{Ga}_2\text{O}_3}{15.74 \text{ g } \text{Ga}_2\text{O}_3} \times 100$

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Resources for Teaching Middle School Science Smithsonian Institution, National Academy of Engineering, National Science Resources Center of the National Academy of Sciences, Institute of Medicine, 1998-04-30 With age appropriate inquiry centered curriculum materials and sound teaching practices middle school science can capture the interest and energy of adolescent students and expand their understanding of the world around them Resources for Teaching Middle School Science developed by the National Science Resources Center NSRC is a valuable tool for identifying and selecting effective science curriculum materials that will engage students in grades 6 through 8 The volume describes more than 400 curriculum titles that are aligned with the National Science Education Standards This completely new guide follows on the success of Resources for Teaching Elementary School Science the first in the NSRC series of annotated guides to hands on inquiry centered curriculum materials and other resources for science teachers The curriculum materials in the new guide are grouped in five chapters by scientific area Physical Science Life Science Environmental Science Earth and Space Science and Multidisciplinary and Applied Science They are also grouped by type core materials supplementary units and science activity books Each annotation of curriculum material includes a recommended grade level a description of the activities involved and of what students can be expected to learn a list of accompanying materials a reading level and ordering information The curriculum materials included in this book were selected by panels of teachers and scientists using evaluation criteria developed for the guide The criteria reflect and incorporate goals and principles of the National Science

Education Standards The annotations designate the specific content standards on which these curriculum pieces focus In addition to the curriculum chapters the guide contains six chapters of diverse resources that are directly relevant to middle school science Among these is a chapter on educational software and multimedia programs chapters on books about science and teaching directories and guides to science trade books and periodicals for teachers and students Another section features institutional resources One chapter lists about 600 science centers museums and zoos where teachers can take middle school students for interactive science experiences Another chapter describes nearly 140 professional associations and U S government agencies that offer resources and assistance Authoritative extensive and thoroughly indexed and the only guide of its kind Resources for Teaching Middle School Science will be the most used book on the shelf for science teachers school administrators teacher trainers science curriculum specialists advocates of hands on science teaching and concerned parents

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Chemistry, 1976 *Chemistry in Canada*, 1956 *The Publishers Weekly*, 1917 *Walford's Guide to Reference Material: Science and technology* Albert John Walford, Library Association, 1999 A revised and updated guide to reference material It contains selective and evaluative entries to guide the enquirer to the best source of reference in each subject area

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Table of Contents Mcgraw Hill Connect Chemistry Study Guide Answers

1. Understanding the eBook Mcgraw Hill Connect Chemistry Study Guide Answers
 - The Rise of Digital Reading Mcgraw Hill Connect Chemistry Study Guide Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Mcgraw Hill Connect Chemistry Study Guide Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Mcgraw Hill Connect Chemistry Study Guide Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Mcgraw Hill Connect Chemistry Study Guide Answers

- Personalized Recommendations
- Mcgraw Hill Connect Chemistry Study Guide Answers User Reviews and Ratings
- Mcgraw Hill Connect Chemistry Study Guide Answers and Bestseller Lists
- 5. Accessing Mcgraw Hill Connect Chemistry Study Guide Answers Free and Paid eBooks
 - Mcgraw Hill Connect Chemistry Study Guide Answers Public Domain eBooks
 - Mcgraw Hill Connect Chemistry Study Guide Answers eBook Subscription Services
 - Mcgraw Hill Connect Chemistry Study Guide Answers Budget-Friendly Options
- 6. Navigating Mcgraw Hill Connect Chemistry Study Guide Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Mcgraw Hill Connect Chemistry Study Guide Answers Compatibility with Devices
 - Mcgraw Hill Connect Chemistry Study Guide Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Mcgraw Hill Connect Chemistry Study Guide Answers
 - Highlighting and Note-Taking Mcgraw Hill Connect Chemistry Study Guide Answers
 - Interactive Elements Mcgraw Hill Connect Chemistry Study Guide Answers
- 8. Staying Engaged with Mcgraw Hill Connect Chemistry Study Guide Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Mcgraw Hill Connect Chemistry Study Guide Answers
- 9. Balancing eBooks and Physical Books Mcgraw Hill Connect Chemistry Study Guide Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Mcgraw Hill Connect Chemistry Study Guide Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Mcgraw Hill Connect Chemistry Study Guide Answers
 - Setting Reading Goals Mcgraw Hill Connect Chemistry Study Guide Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Mcgraw Hill Connect Chemistry Study Guide Answers

- Fact-Checking eBook Content of Mcgraw Hill Connect Chemistry Study Guide Answers
- Distinguishing Credible Sources

13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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