

Question Number	Acceptable Answers	Reject	Mark
4 (f) (i)	$(7.5 \times 0.81) = 6.075 / 6.08 \text{ (g)}$ Ignore sf except 1 sf	6.07 Wrong units	1

Question Number	Acceptable Answers	Reject	Mark
4 (f) (ii)	<p>Look at final answer. 67% scores 3 marks; answers with 3sf rounding to 67 score 2 marks. If this is incorrect follow this scheme:</p> <p>METHOD 1</p> <p>Mol butan-1-ol = $(6.075/74)$ $= 0.0820945$ (1)</p> <p>maximum mass 1-bromobutane = $(0.0820945 \times 137) = 11.246959 \text{ g}$ (1)</p> <p>% yield = $((7.5/11.24659) \times 100)$ $= 66.85)$ $= 67\% \text{ to 2 sf}$ (1)</p> <p>OR METHOD 2</p> <p>$7.5/137 = 0.0547445 \text{ mol (bromobutane)}$ (1)</p> <p>$6.075/74 = 0.0820945 \text{ mol butan-1-ol}$ (1)</p> <p>% yield = $((0.05474455) \times 100 / 0.0820945)$ $= 66.85)$ $= 67\% \text{ to 2 sf}$ (1)</p> <p>Also TE from one step of the calculation to the next and TE on 4f(i) unless yield > 100%.</p> <p>Use of 6.08 gives 0.082161 mol, 11.256216 g bromobutane, final answer 67%</p> <p>11.3g bromobutane gives 66%.</p>	<p>Percentages calculated from volumes with no conversion to mol or mass.</p> <p>$6.075/7.5 \times 100$ $= 81\% \text{ scores 0}$</p> <p>67.0 (This is 3sf)</p>	3

Total for Question 4 = 15 marks

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Table of Contents January 2015 Edexcel Ial Ums

1. Understanding the eBook January 2015 Edexcel Ial Ums
 - The Rise of Digital Reading January 2015 Edexcel Ial Ums
 - Advantages of eBooks Over Traditional Books
2. Identifying January 2015 Edexcel Ial Ums
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an January 2015 Edexcel Ial Ums
 - User-Friendly Interface
4. Exploring eBook Recommendations from January 2015 Edexcel Ial Ums

- Personalized Recommendations
- January 2015 Edexcel Ial Ums User Reviews and Ratings
- January 2015 Edexcel Ial Ums and Bestseller Lists
- 5. Accessing January 2015 Edexcel Ial Ums Free and Paid eBooks
 - January 2015 Edexcel Ial Ums Public Domain eBooks
 - January 2015 Edexcel Ial Ums eBook Subscription Services
 - January 2015 Edexcel Ial Ums Budget-Friendly Options
- 6. Navigating January 2015 Edexcel Ial Ums eBook Formats
 - ePub, PDF, MOBI, and More
 - January 2015 Edexcel Ial Ums Compatibility with Devices
 - January 2015 Edexcel Ial Ums Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of January 2015 Edexcel Ial Ums
 - Highlighting and Note-Taking January 2015 Edexcel Ial Ums
 - Interactive Elements January 2015 Edexcel Ial Ums
- 8. Staying Engaged with January 2015 Edexcel Ial Ums
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers January 2015 Edexcel Ial Ums
- 9. Balancing eBooks and Physical Books January 2015 Edexcel Ial Ums
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection January 2015 Edexcel Ial Ums
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine January 2015 Edexcel Ial Ums
 - Setting Reading Goals January 2015 Edexcel Ial Ums
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of January 2015 Edexcel Ial Ums

- Fact-Checking eBook Content of January 2015 Edexcel Ial Ums
- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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