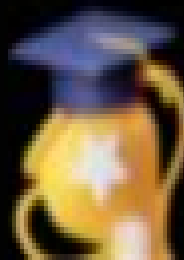
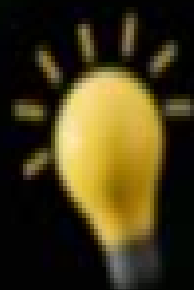
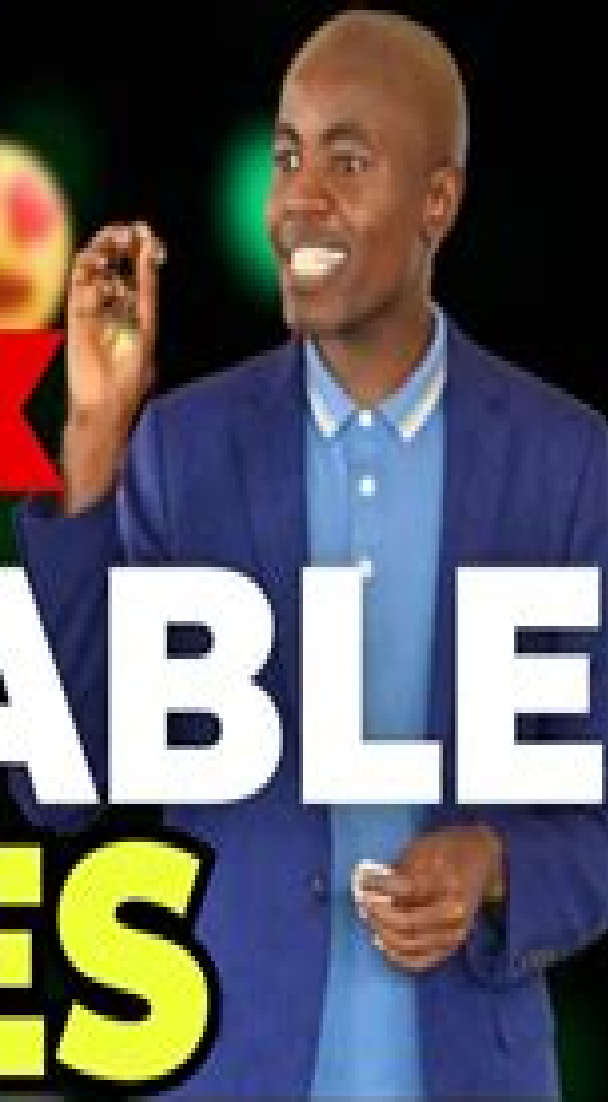


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web since the mole fraction of the  $\text{O}_2$  is  $0.050$  it contributes  $0.050$  or if you will  $5.0\%$  of the total pressure thus total pressure is  $0.21005042 \text{ atm}$  problem 9 a sample of  $1.43 \text{ g}$  of helium and an unweighed quantity of  $\text{O}_2$  are mixed in a flask at room temperature

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