

Limiting Reactant, Theoretical Yield, Actual Yield, & Percent Yield

$22.99 + 35.45 = 58.44$

24



- If we have 53.2 g of Na and 65.8 g of Cl_2 , what are the limiting reactant and theoretical yield?

$$53.2\text{ g Na} \times \frac{1\text{ mol Na}}{22.99\text{ g Na}} \times \frac{2\text{ mol NaCl}}{2\text{ mol Na}} \times \frac{58.44\text{ g NaCl}}{1\text{ mol NaCl}} = 135.23$$

$$65.8\text{ g Cl}_2 \times \frac{1\text{ mol Cl}_2}{70.9\text{ g Cl}_2} \times \frac{2\text{ mol NaCl}}{1\text{ mol Cl}_2} \times \frac{58.44\text{ g NaCl}}{1\text{ mol NaCl}} = 110.5$$

Limiting Reagent And Percent Yield Lab Answers

K Payea



Limiting Reagent And Percent Yield Lab Answers:

Whispering the Secrets of Language: An Psychological Quest through **Limiting Reagent And Percent Yield Lab Answers**

In a digitally-driven earth wherever screens reign great and immediate connection drowns out the subtleties of language, the profound techniques and emotional subtleties concealed within words frequently get unheard. However, situated within the pages of **Limiting Reagent And Percent Yield Lab Answers** a interesting literary value sporting with organic thoughts, lies an exceptional quest waiting to be undertaken. Published by an experienced wordsmith, this charming opus encourages readers on an introspective journey, delicately unraveling the veiled truths and profound affect resonating within ab muscles cloth of every word. Within the emotional depths of the moving review, we can embark upon a genuine exploration of the book is core subjects, dissect its charming publishing design, and succumb to the effective resonance it evokes heavy within the recesses of readers hearts.

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