## Limiting Reagent & Percent Yield Practice Worksheet

- When copper (II) chloride reacts with sodium nitrate, copper (II) nitrate and sodium chloride are formed.
  - a. Write the balanced equation for the reaction given above:

+ NaNO<sub>1</sub> → Cu(NO<sub>1</sub>)<sub>2</sub> + ChrCile CuCl<sub>2</sub> + 2 NaNO<sub>3</sub> → Cu(NO<sub>3</sub>)<sub>2</sub> + 2 NaCl

 If 15 grams of copper (II) chloride react with 20 grams of sodium nitrate, how much sodium chloride can be formed?

1 mol CuCl<sub>2</sub> x 2 mol NaCl x 58.443 g NaCl = 13 g NaCl 15 g CuClax 134.362g CuCl: 1 mol CuCl: 1 mol NaCl

20 g NaNO<sub>3</sub> x 1 mol NaNO<sub>3</sub> x 2 mol NaCl x 58.443 g NaCl = 13.8 g NaCl 84,997g NaNOs 2 mol NaNOs 1 mol NaCl

Only 13 g can be formed

- c. What is the limiting reagent for the reaction in #2? copper (II) chloride
- d. How many grams of copper (II) nitrate is formed?

 $15g CuCl_2 \times 1 \mod CuCl_2 \times 1 \mod Cu(NO_3)_2 \times 187.467 g Cu(NO_3)_2 = 20.9 g Cu(NO_3)_2$ 134,452g CuCl<sub>2</sub> 1 mol CuCl<sub>2</sub> 1 mol Cu(NO<sub>3</sub>)<sub>2</sub>

e. How much of the excess reasent is left over in this reaction?

15g CuCl<sub>2</sub> x 1 mol CuCl<sub>2</sub> x 2 mol NaNO<sub>3</sub> x 84.997g NaNO<sub>3</sub> = 19 g NaNO<sub>3</sub> 134.452g CuCl<sub>2</sub> 1 mol CuCl<sub>2</sub> 1 mol NaNO<sub>3</sub>

20 g NaNOs - 19 g NaNOs = 1 g NaNOs

f. If 11.3 grams of sodium chloride are formed in the reaction described in problem #2, what is Percent yield =  $\frac{\text{actual yield}}{\text{theoretical yield}} \times 100 \text{ percent}$ the percent yield of this reaction?

formed.

11.3/13.0 x 100% = 86.9%

- 2. When lead (II) nitrate reacts with sodium iodide, sodium nitrate and lead (II) iodide are
  - a. Balance the following equation:

Pb(NO<sub>3</sub>)<sub>2 (and</sub> + NaI (and → PbI<sub>2 (a)</sub> + NaNO<sub>3 (and</sub> Pb(NO<sub>3</sub>)<sub>2 tant</sub> + 2 NaI <sub>tant</sub> → PbI<sub>2 tot</sub> + 2 NaNO<sub>3 tant</sub>

 If I start with 25.0 grams of lead (II) nitrate and 15.0 grams of sodium iodide, how many grams of sodium nitrate can be formed?

25.0 g Pb(NO<sub>3</sub>)<sub>2</sub> x 1 mol Pb(NO<sub>3</sub>)<sub>2</sub> x 2 mol NaNO<sub>3</sub> x 84.997g NaNO<sub>3</sub> = 12.8 g NaNO<sub>3</sub> 331.214 g Pb(NO<sub>3</sub>)<sub>2</sub> 1 mol Pb(NO<sub>3</sub>)<sub>2</sub> 1 mol NaNO<sub>3</sub>

15.0 g Nal x 1 mol Nal x 2 mol NaNO<sub>3</sub> x 84.997g NaNO<sub>3</sub> = 8.51 g NaNO<sub>3</sub> 149,89g Nat 2 mol NaI 1 mol NaNOs

8.51 g NaNO3 can be made

# <u>Limiting Reagent And Percent Yield Answer Key 12 3</u>

**Albert A Gayle** 

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## **Table of Contents Limiting Reagent And Percent Yield Answer Key 12 3**

- 1. Understanding the eBook Limiting Reagent And Percent Yield Answer Key 12 3
  - o The Rise of Digital Reading Limiting Reagent And Percent Yield Answer Key 12 3
  - Advantages of eBooks Over Traditional Books
- 2. Identifying Limiting Reagent And Percent Yield Answer Key 12 3
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - o Features to Look for in an Limiting Reagent And Percent Yield Answer Key 12 3
  - User-Friendly Interface
- 4. Exploring eBook Recommendations from Limiting Reagent And Percent Yield Answer Key 12 3
  - Personalized Recommendations
  - $\circ\,$  Limiting Reagent And Percent Yield Answer Key 12 3 User Reviews and Ratings
  - Limiting Reagent And Percent Yield Answer Key 12 3 and Bestseller Lists
- 5. Accessing Limiting Reagent And Percent Yield Answer Key 12 3 Free and Paid eBooks
  - Limiting Reagent And Percent Yield Answer Key 12 3 Public Domain eBooks
  - Limiting Reagent And Percent Yield Answer Key 12 3 eBook Subscription Services
  - Limiting Reagent And Percent Yield Answer Key 12 3 Budget-Friendly Options
- 6. Navigating Limiting Reagent And Percent Yield Answer Key 12 3 eBook Formats

- o ePub, PDF, MOBI, and More
- Limiting Reagent And Percent Yield Answer Key 12 3 Compatibility with Devices
- Limiting Reagent And Percent Yield Answer Key 12 3 Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - o Adjustable Fonts and Text Sizes of Limiting Reagent And Percent Yield Answer Key 12 3
  - Highlighting and Note-Taking Limiting Reagent And Percent Yield Answer Key 12 3
  - o Interactive Elements Limiting Reagent And Percent Yield Answer Key 12 3
- 8. Staying Engaged with Limiting Reagent And Percent Yield Answer Key 12 3
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - $\circ\,$  Following Authors and Publishers Limiting Reagent And Percent Yield Answer Key 12 3
- 9. Balancing eBooks and Physical Books Limiting Reagent And Percent Yield Answer Key 12 3
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Limiting Reagent And Percent Yield Answer Key 12 3
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Limiting Reagent And Percent Yield Answer Key 12 3
  - Setting Reading Goals Limiting Reagent And Percent Yield Answer Key 12 3
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Limiting Reagent And Percent Yield Answer Key 12 3
  - Fact-Checking eBook Content of Limiting Reagent And Percent Yield Answer Key 12 3
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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