



# CHEMICAL BONDS



## IONIC, COVALENT, & METALLIC

**FORMS OF CHEMICAL BONDS**  
Bonding: combination of atoms

Atoms combine to form molecules or compounds by sharing or transferring electrons. This process is called bonding. The resulting molecule or compound has properties different from the individual atoms.

**Types of Bonds:**

- Covalent Bond:** Formed by sharing electrons between two atoms. Example:  $H_2$ ,  $H_2O$ ,  $CO_2$ .
- Ionic Bond:** Formed by the transfer of electrons from one atom to another, creating oppositely charged ions. Example:  $NaCl$ ,  $MgO$ .
- Metallic Bond:** Formed by the delocalization of valence electrons in a metal lattice. Example:  $Fe$ ,  $Cu$ ,  $Al$ .

**Covalent, Ionic, and Metallic Bonds**

Covalent bonds are formed by sharing electrons. They are typically found in nonmetals and their compounds. Ionic bonds are formed by the transfer of electrons. They are typically found in metals and nonmetals. Metallic bonds are formed by the delocalization of electrons. They are typically found in metals.

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## GUIDED READING ACTIVITY

# Ionic Bonds Guided Reading And Study Answers

**Raffaela Di Napoli**



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*Instructor's Manual and Study Guide Answers for the Human Body in Health and Disease* Barbara Janson Cohen, Memmler, 1996

**Study Guide/Selected Solutions Manual** Julie R. Frentrup, 2002-06 Contains a brief overview of every chapter review of skills self tests and the answers and detailed solutions to all odd numbered end of chapter problems in the text book

*Chemical Interactions*, 2005

**Prentice Hall Science Explorer** Michael J. Padilla, 2002      **Glencoe Science** McGraw-Hill Staff, 2001-06      Study

[Guide](#) George Karleskint,1991      **Addison-Wesley Chemistry** Antony C. Wilbraham,Dennis D. Staley,Michael S. Matta,Edward L. Waterman,Prentice-Hall Staff,2001-02 To purchase or download a workbook click on the Purchase or Download button to the left To purchase a workbook enter the desired quantity and click Add to Cart To download a free workbook right click the FREE Download PDF link and save to your computer This will result in a faster download as opposed to left clicking and opening the link      [Teacher's Manual and Resource Guide for Exploring the Sciences](#) Herbert Drapkin,1964      [Focus on physical science](#) Charles H. Heimler,1989      **Holt Physical Science** Mapi M. Cuevas,1994  
**Biochemistry Explained** Thomas Millar,2018-12-14 Biochemistry Explained employs an innovative approach which has proven highly successful in the author s own classes The author establishes a thorough understanding of the foundations of and common linkages between molecular structures and reactions so that eventual interpretation of complex biochemical pathways and reactions is easy All of the major molecular structures and biochemical pathways are explained and for the most part these center on mammalian biochemistry The text is supported by biochemical nomenclature and questions to bear in mind while reading Higher learning sections are also provided for advanced students Written in an informal conversational style this textbook will serve as an invaluable resource for any student who is struggling with the standard texts and for postgraduate students who need to refresh their knowledge      [Study and Problem Solving Guide to Accompany Principles of Modern Chemistry, Oxtoby/Nachtrieb](#) Wade A. Freeman,1986      **Prentice Hall Chemistry** ,2000      **GED Science** Cambridge,1993-11      *Programmed Learning and Individually Paced Instruction* Carl H. Hendershot,1973      **Books in Print Supplement** ,2002      **The English Catalogue of Books** ,1884      [Chemistry of Matter](#) ,1997

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