

### Ions

How are ions made from neutral atoms?

You have learned that not all atoms of an element are the same. Variation in the number of neutrons results in different isotopes of the element. In this activity we will explore another variation that can take place—the loss and gain of electrons. The exchange of electrons between atoms is a very common way for chemical change to take place. We will see it many times throughout the year.

1. Use Model 1 (on the next page) to complete the following table.

	Metal or Nonmetal?	Is the number of protons the same in the atom and the ion?	Is the number of neutrons the same in the atom and the ion?	Is the number of electrons the same in the atom and the ion?	Charge on the ion
Lithium	Metal	Yes	Yes	No	1+
Magnesium	Metal	Yes	Yes	No	2+
Aluminum	Metal	Yes	Yes	No	3+
Fluorine	Non-Metal	Yes	Yes	No	1-
Oxygen	nonmetal	Yes	Yes	No	2-
Nitrogen	non-metal	Yes	Yes	No	3-

2. Based on the table, what distinguishes a neutral atom from an ion?  
 a neutral atom has the same amount of  $P^+$  and  $E^-$
3. Examine the isotope symbols in Model 1.
- a) Where is the ion charge located in the isotope symbol?  
 the charge of an ion comes after the charge number
- b) Is a charge indicated on the neutral atoms? If yes, where is it located?  
 a charge isn't indicated in neutral atoms b/c they don't have a charge and  $P^+ - E^- = 0$ .
4. Which subatomic particle carries a positive charge?  
 $E^-$ ,  $Mg$ ,  $Al$
5. Which subatomic particle carries a negative charge?  
 $F$ ,  $O$ ,  $N$
6. Propose a mathematical equation to calculate the charge on an ion from the number of protons and electrons in an ion. Confirm that your equation works using a positive ion example and a negative ion example from Model 1.

$$\begin{aligned} \text{Charge of an ion} &= \text{number of protons} - \text{number of electrons} \\ \text{Nitrogen} &= P^+ - E^- \\ &= 7 - 10 \\ &= -3 (3-) \end{aligned}$$

$$\begin{aligned} \text{Aluminum} &= P^+ - E^- \\ &= 13 - 10 \\ &= +3 (3+) \end{aligned}$$

# Ions Pogil Answer Key Chemistry

**Rachel Turoscy**



## **Ions Pogil Answer Key Chemistry:**

Analytical Chemistry Juliette Lantz, Renée Cole, The POGIL Project, 2014-12-31 An essential guide to inquiry approach instrumental analysis Analytical Chemistry offers an essential guide to inquiry approach instrumental analysis collection The book focuses on more in depth coverage and information about an inquiry approach This authoritative guide reviews the basic principles and techniques Topics covered include method of standard the microscopic view of electrochemistry calculating cell potentials the BerriLambert atomic and molecular absorption processes vibrational modes mass spectra interpretation and much more

Empowering Science and Mathematics for Global Competitiveness Yuli Rahmawati, Peter Charles Taylor, 2019-06-07 This conference proceedings focuses on enabling science and mathematics practitioners and citizens to respond to the pressing challenges of global competitiveness and sustainable development by transforming research and teaching of science and mathematics The proceedings consist of 82 papers presented at the Science and Mathematics International Conference SMIC 2018 organised by the Faculty of Mathematics and Natural Sciences Universitas Negeri Jakarta Indonesia The proceedings are organised in four parts Science Science Education Mathematics and Mathematics Education The papers contribute to our understanding of important contemporary issues in science especially nanotechnology materials and environmental science science education in particular environmental sustainability STEM and STEAM education 21st century skills technology education and green chemistry and mathematics and its application in statistics computer science and mathematics education

*Doklady*, 1996

**Reactions of Ions** William Edward Marsden, 1996

**The Answer Key: A Comprehensive Explanation of Problem Solving Methods for General Chemistry Success (Volume Two) (First Edition)** Rachel Turoscy, 2018-08-09 The Answer Key A Comprehensive Explanation of Problem Solving Methods for General Chemistry Success Volume 2 is a concise and accessible textbook that covers the critical information a student needs to understand the basic mathematics used in chemistry courses The book provides easy to understand step by step instructions for solving general chemistry problems The book begins with chapters dedicated to solutions kinetics and liquids solids and phase changes In subsequent chapters the text covers important topics like equilibrium concentrations strong and weak acids and bases the Common Ion Effect and reaction mechanisms It also covers the equilibrium between a solid and its respective ions in a solution as well as the second law of Thermodynamics The text also addresses Gibbs Free Energy equilibrium constants and electrolysis calculations Each chapter contains sample problems and practice problems to help further understanding of how math and chemistry go hand in hand The Answer Key is an excellent resource for any undergraduate course that deals with the basic concepts of general chemistry

**Selected Chemistry** Topich, John E. McMurry, 1997-12 Contains solutions to all in chapter problems all understanding key concept questions and selected end of chapter problems

**Science Foundation Course**, 1971

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**Fundamentals of Introductory Chemistry** Darrell D. Ebbing, 1998-06-01 *Practical/Laboratory Manual Chemistry Class - XI* Er. Meera Goyal, 2021-05-29

1 Basic Laboratory Techniques

- 1 To cut a glass tube or glass rod
- 2 To bend the glass rod at an angle
- 3 To draw a glass jet from a glass tube
- 4 To bore a cork and fit a glass tube into it

Viva Voce

2 Characterisation and Purification of Chemical Substances

- 1 To determine the melting point of the given unknown organic compound and its identification simple laboratory technique

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- 2 To determine the boiling point of a given liquid when available in small quantity simple laboratory method

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- 3 To prepare crystals of pure potash alum  $K_2SO_4 \cdot Al_2(SO_4)_3 \cdot 24H_2O$  from the given impure sample
- 4 To prepare the pure crystals of copper sulphate from the given crude sample
- 5 To prepare pure crystals of benzoic acid from a given impure sample

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3 Measurement of pH Values

- 1 To determine the pH value of vegetable juices fruit juices tap water and washing soda by using universal pH paper
- 2 To determine and compare the pH values of solutions of strong acid HCl and weak acid  $CH_3COOH$  of same concentration
- 3 To study the pH change in the titration of strong base Vs strong acid by using universal indicator paper
- 4 To study the pH

change by common ion  $\text{CH}_3\text{COO}^-$  ion in case of weak acid  $\text{CH}_3\text{COOH}$  5 To determine the change in pH value of weak base  $\text{NH}_4\text{OH}$  in presence of a common ion  $\text{NH}_4^+$  Viva Voce 4 Chemical Equilibrium 1 To study the shift in equilibrium between ferric ions and thiocyanate ions by changing the concentrations of either of the ions 2 To study the shift in equilibrium between  $\text{Co}^{2+}$  and  $\text{Cl}^-$  ions by changing the concentrations of either of the ions Viva Voce 5 Quantitative Analysis 1 To prepare M 10 oxalic acid solution by direct weighing method 2 To prepare M 10 solution of sodium carbonate by direct weighing method 3 To determine the strength of given solution of sodium hydroxide by titrating it against N 10 or M 20 solution of oxalic acid 4 To determine the strength of a given solution of hydrochloric acid by titrating it against a standard N 10 or M 20 sodium carbonate solution Viva Voce 6 Qualitative Analysis Analysis of Anions Analysis of Cations Viva Voce 7 Detection of Elements in Organic Compounds 1 To detect the presence of nitrogen sulphur and halogens in a given organic compound by Lassaigne's test 2 To detect the presence of nitrogen sulphur and halogens in the given organic compound sample number by Lassaigne's test Viva Voce INVESTIGATORY PROJECTS 1 Checking of Bacterial Contamination in Water 1 To check the bacterial contamination in drinking water by testing sulphide ions Viva Voce 2 Methods of Water Purification 1 To purify water from suspended impurities by using sedimentation 2 To purify water by boiling 3 To purify water by distillation method 4 To purify water by reverse osmosis technique 5 To purify water by GAC method 6 To purify water by bleach treatment 7 To purify water by oxidising agent 8 To purify water by ozone treatment method Viva Voce 3 Water Analysis 1 To test the hardness of different water samples Viva Voce 4 Foaming Capacity of Various Soaps 1 To compare the foaming capacity of different washing soaps 2 To study the effect of addition of sodium carbonate on foaming capacity of washing soap Viva Voce 5 Tea Analysis 1 To study the acidity of different samples of tea leaves tea by using pH paper Viva Voce 6 Analysis of Fruits and Vegetable Juices 1 To analysis the fruit and vegetable juices for the constituent present in them Viva Voce 7 Rate of Evaporation 1 To study the rate of evaporation of different liquids Viva Voce 8 Effect of Acids and Bases on Tensile Strength of Fibres 1 To compare the tensile strength of natural fibres and synthetic fibres 2 To study the effect of acids and bases on tensile strength of different fibres Viva Voce

Fundamentals of Introductory Chemistry Darrell D. Ebbing, 1998-06-01

**Syllabus for Chemistry 130-230** University of Chicago, 1935

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*Ions in Solution. Prep. by the Science Foundation Course Team (of The) Open University (Bletchley) ,1971*

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form for ease of understanding and systematic learning. Clearly illustrated diagrams and tables are also included to help students understand difficult concepts and principles. The author believes that students will find this book a good source of relevant and important notes and a useful revision guide and study aid.

**Chemistry: Special Study: Ion Exchange**  
Nuffield Foundation. Science Teaching Project, 1970

**Chemistry** Richard S. Moog, John J. Farrell, 2017-06-26

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