

POGIL 2 Ions

How are ions made from neutral atoms?

You have learned that not all atoms of an element are the same. Variation in the number of neutrons results in different isotopes of the element. In this activity we will explore another variation that can take place—the loss and gain of electrons. The exchange of electrons between atoms is a very common way for chemical change to take place. We will see it many times throughout the year.

1. Use Diagram 1 on the next page to complete the following table.

| | Metal or Nonmetal | Number of protons in the atom | Number of protons in the ion | Number of neutrons in the atom | Number of neutrons in the ion | Number of electrons in the atom | Number of electrons in the ion | Charge on the ion |
|-----------|-------------------|-------------------------------|------------------------------|--------------------------------|-------------------------------|---------------------------------|--------------------------------|-------------------|
| Lithium | metal | 3 | 3 | 4 | 4 | 3 | 2 | 1+ |
| Magnesium | <i>Metal</i> | 12 | 12 | 12 | 12 | 12 | 10 | 2+ |
| Aluminum | <i>Metalloid</i> | 13 | 13 | 14 | 14 | 13 | 10 | 3+ |
| Fluorine | <i>Non metal</i> | 9 | 9 | 10 | 10 | 9 | 10 | 1- |
| Oxygen | nonmetal | 8 | 8 | 8 | 8 | 8 | 10 | 2- |
| Nitrogen | <i>Non metal</i> | 7 | 7 | 7 | 7 | 7 | 10 | 3- |

2. Based on the table you completed in Question 1, what distinguishes a neutral atom from an ion?

A neutral atom has the same number of protons and electrons (no charge). An ion has a charge (# of protons does not equal # of electrons).

3. Examine the isotope symbols in Diagram 1.

- a. Where is the ion charge located in the isotope symbol?

Top right corner

- b. Is a charge indicated on the neutral atoms? If yes, where is it located?

No

4. Which subatomic particle carries a positive charge?

Protons are positive

5. Which subatomic particle carries a negative charge?

Electrons are negative

6. Propose a mathematical equation to calculate the charge on an ion from the number of protons and electrons in an ion. Confirm that your equation works using two positive ion examples and two negative ion examples from Diagram 1.

The charge of an ion = # of protons - # of electrons

Ions Pogil Packet Answer Key

**Juliette Lantz, Renée Cole, The POGIL
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Table of Contents Ions Pogil Packet Answer Key

1. Understanding the eBook Ions Pogil Packet Answer Key
 - The Rise of Digital Reading Ions Pogil Packet Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Ions Pogil Packet Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Ions Pogil Packet Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Ions Pogil Packet Answer Key
 - Personalized Recommendations
 - Ions Pogil Packet Answer Key User Reviews and Ratings
 - Ions Pogil Packet Answer Key and Bestseller Lists
5. Accessing Ions Pogil Packet Answer Key Free and Paid eBooks
 - Ions Pogil Packet Answer Key Public Domain eBooks
 - Ions Pogil Packet Answer Key eBook Subscription Services
 - Ions Pogil Packet Answer Key Budget-Friendly Options
6. Navigating Ions Pogil Packet Answer Key eBook Formats
 - ePub, PDF, MOBI, and More
 - Ions Pogil Packet Answer Key Compatibility with Devices
 - Ions Pogil Packet Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Ions Pogil Packet Answer Key
 - Highlighting and Note-Taking Ions Pogil Packet Answer Key
 - Interactive Elements Ions Pogil Packet Answer Key
8. Staying Engaged with Ions Pogil Packet Answer Key

- Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Ions Pogil Packet Answer Key
9. Balancing eBooks and Physical Books Ions Pogil Packet Answer Key
- Benefits of a Digital Library
 - Creating a Diverse Reading Collection Ions Pogil Packet Answer Key
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Ions Pogil Packet Answer Key
- Setting Reading Goals Ions Pogil Packet Answer Key
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Ions Pogil Packet Answer Key
- Fact-Checking eBook Content of Ions Pogil Packet Answer Key
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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