

Chapter 12 DNA and RNA**Section 12–1 DNA (pages 287–294)**

This section tells about the experiments that helped scientists discover the relationship between genes and DNA. It also describes the chemical structure of the DNA molecule.

Griffith and Transformation (pages 287–289)

1. What did Frederick Griffith want to learn about bacteria? He wanted to learn how certain types of bacteria produce pneumonia.
2. The strain of bacteria that caused pneumonia grew into smooth colonies on culture plates; harmless bacteria produced colonies with rough edges.
3. Circle the letter of each sentence that is true about Griffith's experiment.
 - ☒ a. Mice injected with bacteria from smooth colonies died.
 - ☐ b. Mice injected with bacteria from rough colonies died.
 - ☐ c. Mice injected with heat-killed bacteria from smooth colonies died.
 - ☒ d. Mice injected with a mixture of bacteria from heat-killed smooth colonies and live rough colonies died.
4. What result from Griffith's experiment suggested that the cause of pneumonia was not a chemical poison released by the disease-causing bacteria? The mice survived after being injected with heat-killed disease-causing bacteria.
5. What is transformation? It is the process by which one strain of bacteria changes into another.
6. What hypothesis did Griffith form from the results of his experiments? Some factor, which might be a gene, was transferred from the heat-killed bacteria cells into the live cells.

Avery and DNA (page 289)

7. Is the following sentence true or false? Avery and his colleagues thought that the molecule required in transformation might also be the molecule of the gene.
true
8. Briefly describe how Avery and his group determined which molecule was most important for transformation? They treated the extract of heat-killed bacteria with enzymes that destroyed proteins, lipids, carbohydrates, and other molecules, including RNA and DNA.

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Prentice Hall Physical Science Concepts in Action Program Planner National Chemistry Physics Earth Science, 2003-11 Prentice Hall Physical Science Concepts in Action helps students make the important connection between the science they read and what they experience every day Relevant content lively explorations and a wealth of hands on activities take students understanding of science beyond the page and into the world around them Now includes even more technology tools and activities to support differentiated instruction **Prentice Hall Chemistry**, 2000 **Study Guide,**

Introductory Chemistry, a Foundation ; Introductory Chemistry ; Basic Chemistry Iris Stovall, Donald J. DeCoste, 2000

Physical Inorganic Chemistry S. F. A. Kettle, 2013-11-11 GEORGE CHRISTOU Indiana University Bloomington I am no doubt representative of a large number of current inorganic chemists in having obtained my undergraduate and postgraduate degrees in the 1970s It was during this period that I began my continuing love affair with this subject and the fact that it happened while I was a student in an organic laboratory is beside the point I was always enchanted by the more physical aspects of inorganic chemistry while being captivated from an early stage by the synthetic side and the measure of creation with a small c that it entails I nevertheless found the application of various theoretical spectroscopic and physicochemical techniques to inorganic compounds to be fascinating stimulating educational and downright exciting The various bonding theories for example and their use to explain or interpret spectroscopic observations were more or less universally accepted as belonging within the realm of inorganic chemistry and textbooks of the day had whole sections on bonding theories magnetism kinetics electron transfer mechanisms and so on However things changed and subsequent inorganic chemistry teaching texts tended to emphasize the more synthetic and descriptive side of the field There are a number of reasons for this and they no doubt include the rise of diamagnetic organometallic chemistry as the dominant subdiscipline within inorganic chemistry and its relative narrowness vis d vis physical methods required for its prosecution Addison-Wesley Chemistry Antony C. Wilbraham, Dennis D. Staley, Michael S. Matta, Edward L. Waterman, Prentice-Hall Staff, 2001-02 To purchase or download a workbook click on the Purchase or Download button to the left To purchase a workbook enter the desired quantity and click Add to Cart To download a free workbook right click the FREE Download PDF link and save to your computer This will result in a faster download as opposed to left clicking and opening the link **Books in Print**

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