2 Some metal carbonates cannot be obtained in a pure state. For example magnesium carbonate exists in a 'basic' form, in which magnesium hydroxide is also present.

One possible chemical formula of basic magnesium carbonate is MgCO₃.Mg(OH)₃.2H₃O.

When basic magnesium carbonate is heated, if the possible formula were correct, it would decompose as shown below.

$$MgCO_3.Mg(OH)_2.2H_2O(s) \rightarrow 2MgO(s) + CO_2(g) + 3H_2O(g)$$

In this experiment, you will decompose basic magnesium carbonate by heating it, and you will use your results to determine whether this possible formula is correct.

FB 4 is basic magnesium carbonate.

(a) Method

Read through the method before starting any practical work and prepare a table for your results in the space below.

- Weigh a crucible with its lid and record the mass.
- Add 1.1-1.3g of FB 4 to the crucible. Weigh the crucible and lid with FB 4 and record the mass.
- Place the crucible on the pipe-clay triangle and remove the lid.
- Heat the crucible and contents gently for about one minute.
- Then heat the crucible and contents strongly for about four minutes.
- Replace the lid and allow the crucible to cool for at least five minutes.
- While the crucible is cooling, you may wish to begin work on Question 3.
- Re-weigh the crucible and contents with lid. Record the mass.
- Calculate, and record, the mass of FB 4 used and the mass of residue obtained.

<u>Isa Paper 2 Chemistry Metal Carbonates Ore</u>

Richard Bailey

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