

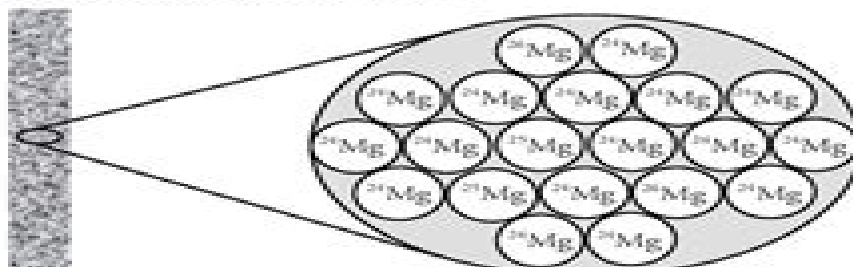
Average Atomic Mass

How are the masses on the periodic table determined?

Why?

Most elements have more than one naturally occurring isotope. As you learned previously, the atoms of those isotopes have the same atomic number (number of protons), making them belong to the same element, but they have different mass numbers (total number of protons and neutrons) giving them different atomic masses. So which mass is put on the periodic table for each element? Is it the most common isotope's mass? The heaviest mass? This activity will help answer that question.

Model 1 – A Strip of Magnesium Metal



1. Write in the atomic number for each Mg atom in Model 1.

All are 12

2. What are the mass numbers of the naturally occurring isotopes of magnesium shown in Model 1?

24, 25, 26

3. Do all of the atoms of magnesium in Model 1 have the same atomic mass? Explain.

No there are 3 different mass numbers which mean there are 3 different masses

4. For the sample of 20 atoms of magnesium shown in Model 1, draw a table indicating the mass numbers of the three isotopes and the number of atoms of each isotope present.

Mass #	# of atoms
24	16
25	2
26	2

5. Which isotope of magnesium is the most common in Model 1?

Magnesium 24 is the most common

6. Based on Model 1 and the table you created in Question 4, for every 10 atoms of magnesium, approximately how many atoms of each isotope will be found?

Mg 24=8 Mg 25=1 Mg 26=1

Isotopes And Average Atomic Mass Answer Key

Melvin D. Joesten



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