

## Titration of Oxalic Acid against Sodium Hydroxide

**Aim:** To find out how strength of a sodium hydroxide solution is by titrating it with a standard solution of (N/10) oxalic acid.

### Theory

This experiment involves titrating a weak acid, oxalic acid, with a strong base, sodium hydroxide, using phenolphthalein as an indicator. The reaction between oxalic acid and sodium hydroxide is:



Because sodium hydroxide isn't a primary standard, we prepared a standard solution of oxalic acid for standardizing sodium hydroxide.

In an acid-base titration, at the endpoint, the amount of acid becomes equal to the amount of base, resulting in a neutral solution.

### Materials

- Burette
- Pipette
- Conical flask
- Burette stand
- Funnel
- Stirrer
- White glazed tile
- Measuring flask
- Oxalic acid (solid)
- Oxalic acid solution
- Sodium hydroxide solution
- Phenolphthalein indicator

### Procedure

#### (a) Preparing a 0.1N Standard Oxalic Acid Solution:

1. To begin, take a watch glass and clean it with distilled water, ensuring it is completely dry. Then, weigh the watch glass accurately and record its weight in your notebook.
2. Next, measure 6.30 grams of oxalic acid carefully and place it on the watch glass. Record this weight in your notebook as well.
3. Using a funnel, gently transfer the oxalic acid from the watch glass into a clean and dry measuring flask. To remove any remaining particles on the watch glass, rinse it with distilled water, being cautious not to exceed 50 mL of water.

# Lab 57 Titration Oxalic Acid

**S Ashworth**

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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