

- c. Select OK. The graph should indicate a direct relationship between absorbance and concentration, (Beer's law). The regression line should closely fit the five data points and pass through (or near) the origin of the graph. Record the linear fit equation in your data table.

Part II Test an Unknown Solution of SCN^-

10. Obtain about 10 mL of the unknown SCN^- solution. Record its number. Use a pipet to measure out 5.0 mL of the unknown into a clean and dry 100 mL beaker. Add precisely 5.0 mL of 0.200 M $\text{Fe}(\text{NO}_3)_3$ and 40.0 mL of distilled water to the beaker. Stir the mixture thoroughly.
11. Using the solution in the beaker, rinse a cuvette twice with ~1 mL amounts and then fill it 3/4 full. Place the cuvette of unknown in the Spectrovis.
12. Determine the concentration of the unknown SCN^- solution.
 - a. Tap the Meter tab.
 - b. Monitor the absorbance value. When this value has stabilized, record it in your data table.
 - c. Use the equation of a straight line that you determined from your test solutions to determine the concentration of FeSCN^{2+} of your unknown. (Note: that y = absorbance and x = concentration, you are solving for " x ")

Part III Prepare and Test Equilibrium Systems

13. Prepare four test tubes of solutions, according to the chart below. Repeat Steps 11 and 12 from Part II to test the absorbance values of each mixture. Record the test results in your data table. Record the $[\text{FeSCN}^{2+}]$ in the table in Data Analysis section #3. **Note:** You are using 0.0020 M $\text{Fe}(\text{NO}_3)_3$ in this test. Note that the blank solution is different.

Test tube number	0.0020 M $\text{Fe}(\text{NO}_3)_3$ (mL)	0.0020 M SCN^- (mL)	H_2O (mL)
1	3.00	2.00	5.00
2	3.00	3.00	4.00
3	3.00	4.00	3.00
4	3.00	5.00	2.00
blank	3.00	0.00	7.00

blank { 3 mL Iron (III) nitrate
 7 mL water
 should be clear

Lab 11 Rate And Equilibrium Report Answer

Society of Chemical Industry (Great Britain)



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- Chapter 5: Conclusion

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