

WORKSHEET: BALANCING EQUATIONS

Balance the following equations by placing small whole-number coefficients in the appropriate places.

Synthesis

1. $\underline{\hspace{1cm}} \text{Ba} + \underline{\hspace{1cm}} \text{O}_2 \rightarrow \underline{\hspace{1cm}} \text{BaO}$
2. $\underline{\hspace{1cm}} \text{C} + \underline{\hspace{1cm}} \text{S} \rightarrow \underline{\hspace{1cm}} \text{CS}_2$
3. $\underline{\hspace{1cm}} \text{Li} + \underline{\hspace{1cm}} \text{O}_2 \rightarrow \underline{\hspace{1cm}} \text{Li}_2\text{O}$
4. $\underline{\hspace{1cm}} \text{Mg} + \underline{\hspace{1cm}} \text{N}_2 \rightarrow \underline{\hspace{1cm}} \text{Mg}_3\text{N}_2$
5. $\underline{\hspace{1cm}} \text{FeCl}_2 + \underline{\hspace{1cm}} \text{Cl}_2 \rightarrow \underline{\hspace{1cm}} \text{FeCl}_3$

Decomposition or Analysis

6. $\underline{\hspace{1cm}} \text{KClO}_3 \rightarrow \underline{\hspace{1cm}} \text{KCl} + \underline{\hspace{1cm}} \text{O}_2$
7. $\underline{\hspace{1cm}} \text{Ag}_2\text{O} \rightarrow \underline{\hspace{1cm}} \text{Ag} + \underline{\hspace{1cm}} \text{O}_2$
8. $\underline{\hspace{1cm}} \text{CuCO}_3 \rightarrow \underline{\hspace{1cm}} \text{CuO} + \underline{\hspace{1cm}} \text{CO}_2$
9. $\underline{\hspace{1cm}} \text{AuBr}_3 \rightarrow \underline{\hspace{1cm}} \text{Au} + \underline{\hspace{1cm}} \text{Br}_2$
10. $\underline{\hspace{1cm}} \text{UF}_4 \rightarrow \underline{\hspace{1cm}} \text{U} + \underline{\hspace{1cm}} \text{F}_2$

Single Replacement

11. $\underline{\hspace{1cm}} \text{Zn} + \underline{\hspace{1cm}} \text{HCl} \rightarrow \underline{\hspace{1cm}} \text{ZnCl}_2 + \underline{\hspace{1cm}} \text{H}_2$
12. $\underline{\hspace{1cm}} \text{Zn} + \underline{\hspace{1cm}} \text{CuSO}_4 \rightarrow \underline{\hspace{1cm}} \text{ZnSO}_4 + \underline{\hspace{1cm}} \text{Cu}$
13. $\underline{\hspace{1cm}} \text{Cu} + \underline{\hspace{1cm}} \text{AgNO}_3 \rightarrow \underline{\hspace{1cm}} \text{Cu}(\text{NO}_3)_2 + \underline{\hspace{1cm}} \text{Ag}$
14. $\underline{\hspace{1cm}} \text{K} + \underline{\hspace{1cm}} \text{H}_2\text{O} \rightarrow \underline{\hspace{1cm}} \text{KOH} + \underline{\hspace{1cm}} \text{H}_2$
15. $\underline{\hspace{1cm}} \text{Al} + \underline{\hspace{1cm}} \text{CuCl}_2 \rightarrow \underline{\hspace{1cm}} \text{AlCl}_3 + \underline{\hspace{1cm}} \text{Cu}$

Double Replacement

16. $\underline{\hspace{1cm}} \text{BaCl}_2 + \underline{\hspace{1cm}} \text{Na}_2\text{SO}_4 \rightarrow \underline{\hspace{1cm}} \text{NaCl} + \underline{\hspace{1cm}} \text{BaSO}_4$
17. $\underline{\hspace{1cm}} \text{ZnCl}_2 + \underline{\hspace{1cm}} (\text{NH}_4)_2\text{S} \rightarrow \underline{\hspace{1cm}} \text{NH}_4\text{Cl} + \underline{\hspace{1cm}} \text{ZnS}$
18. $\underline{\hspace{1cm}} \text{NaOH} + \underline{\hspace{1cm}} \text{HCl} \rightarrow \underline{\hspace{1cm}} \text{NaCl} + \underline{\hspace{1cm}} \text{H}_2\text{O}$
19. $\underline{\hspace{1cm}} \text{FeS} + \underline{\hspace{1cm}} \text{HCl} \rightarrow \underline{\hspace{1cm}} \text{FeCl}_2 + \underline{\hspace{1cm}} \text{H}_2\text{S}$
20. $\underline{\hspace{1cm}} \text{AlCl}_3 + \underline{\hspace{1cm}} \text{NaOH} \rightarrow \underline{\hspace{1cm}} \text{Al}(\text{OH})_3 + \underline{\hspace{1cm}} \text{NaCl}$

Combustion

21. $\underline{\hspace{1cm}} \text{CH}_4 + \underline{\hspace{1cm}} \text{O}_2 \rightarrow \underline{\hspace{1cm}} \text{CO}_2 + \underline{\hspace{1cm}} \text{H}_2\text{O}$
22. $\underline{\hspace{1cm}} \text{C}_2\text{H}_6 + \underline{\hspace{1cm}} \text{O}_2 \rightarrow \underline{\hspace{1cm}} \text{CO}_2 + \underline{\hspace{1cm}} \text{H}_2\text{O}$
23. $\underline{\hspace{1cm}} \text{C}_3\text{H}_8 + \underline{\hspace{1cm}} \text{O}_2 \rightarrow \underline{\hspace{1cm}} \text{CO}_2 + \underline{\hspace{1cm}} \text{H}_2\text{O}$
24. $\underline{\hspace{1cm}} \text{C}_4\text{H}_{10} + \underline{\hspace{1cm}} \text{O}_2 \rightarrow \underline{\hspace{1cm}} \text{CO}_2 + \underline{\hspace{1cm}} \text{H}_2\text{O}$
25. $\underline{\hspace{1cm}} \text{C}_5\text{H}_{12} + \underline{\hspace{1cm}} \text{O}_2 \rightarrow \underline{\hspace{1cm}} \text{CO}_2 + \underline{\hspace{1cm}} \text{H}_2\text{O}$

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Table of Contents Names And Formulas Worksheet Physical Science 233

1. Understanding the eBook Names And Formulas Worksheet Physical Science 233
 - The Rise of Digital Reading Names And Formulas Worksheet Physical Science 233
 - Advantages of eBooks Over Traditional Books
2. Identifying Names And Formulas Worksheet Physical Science 233
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Names And Formulas Worksheet Physical Science 233
 - User-Friendly Interface
4. Exploring eBook Recommendations from Names And Formulas Worksheet Physical Science 233
 - Personalized Recommendations
 - Names And Formulas Worksheet Physical Science 233 User Reviews and Ratings

- Names And Formulas Worksheet Physical Science 233 and Bestseller Lists
- 5. Accessing Names And Formulas Worksheet Physical Science 233 Free and Paid eBooks
 - Names And Formulas Worksheet Physical Science 233 Public Domain eBooks
 - Names And Formulas Worksheet Physical Science 233 eBook Subscription Services
 - Names And Formulas Worksheet Physical Science 233 Budget-Friendly Options
- 6. Navigating Names And Formulas Worksheet Physical Science 233 eBook Formats
 - ePub, PDF, MOBI, and More
 - Names And Formulas Worksheet Physical Science 233 Compatibility with Devices
 - Names And Formulas Worksheet Physical Science 233 Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Names And Formulas Worksheet Physical Science 233
 - Highlighting and Note-Taking Names And Formulas Worksheet Physical Science 233
 - Interactive Elements Names And Formulas Worksheet Physical Science 233
- 8. Staying Engaged with Names And Formulas Worksheet Physical Science 233
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Names And Formulas Worksheet Physical Science 233
- 9. Balancing eBooks and Physical Books Names And Formulas Worksheet Physical Science 233
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Names And Formulas Worksheet Physical Science 233
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Names And Formulas Worksheet Physical Science 233
 - Setting Reading Goals Names And Formulas Worksheet Physical Science 233
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Names And Formulas Worksheet Physical Science 233
 - Fact-Checking eBook Content of Names And Formulas Worksheet Physical Science 233
 - Distinguishing Credible Sources

13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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