

Naming Molecular Compounds

Chem Worksheet 9-2

Name _____

A **molecular compound** is a group of atoms held together by a covalent bond. Compounds made entirely of non-metals are generally molecular compounds. Carbon tetrachloride, CCl_4 , is an example of a molecular compound. When naming these compounds prefixes are used to denote how many of each atom is bonded in the compound. However, the prefix *mono-* is not used with the first element in the compound, even if there is only one element. The ending of the second element in the compound is always changed to *-ide*, in the same way the ending is changed for monatomic anions.

Rules for naming Molecular Compounds

1. Name the first element using the element's full name.
2. Name the second element using the *-ide* ending.
3. Use prefixes to tell how many of each element is present. (do not use the prefix *mono-* on the first element).

Naming Prefixes

1	mono-
2	di-
3	tri-
4	tetra-
5	penta-
6	hexa-
7	hepta-
8	octa-
9	nona-
10	deca-

Examples

#1. Write the chemical formula for diphosphorus pentoxide

- this compound contains two phosphorus atoms and five oxygen atoms



#2. Name the following compound: IF_7

- there is one iodine and there are seven fluorine atoms.

Iodine heptafluoride

(the prefix *mono-* is not used on the first element and that the ending of fluorine is changed to *-ide*.)

Fill in the following table with the missing information.

	Formula	Name
1.	SO_2	
2.		Sulfur trioxide
3.	N_2O_5	
4.		Chlorine dioxide
5.	P_2O_5	
6.		Carbon disulfide
7.	NO_2	
8.	N_2Cl_4	
9.		Xenon difluoride
10.	S_2Cl_2	
11.		Iodine trichloride
12.	P_2S_5	

	Formula	Name
13.	SF_6	
14.		Tetraphosphorus hexasulfide
15.	SeO_2	
16.		Ammonia
17.		Boron trichloride
18.	N_2O	
19.	BrF_3	
20.		Carbon dioxide
21.		Carbon monoxide
22.	ClF_3	
23.		Iodine monochloride
24.	CH_4	

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CH Cherryholmes



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Table of Contents Naming Molecular Compounds Pogil Answers

1. Understanding the eBook Naming Molecular Compounds Pogil Answers
 - The Rise of Digital Reading Naming Molecular Compounds Pogil Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Naming Molecular Compounds Pogil Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Naming Molecular Compounds Pogil Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Naming Molecular Compounds Pogil Answers
 - Personalized Recommendations
 - Naming Molecular Compounds Pogil Answers User Reviews and Ratings
 - Naming Molecular Compounds Pogil Answers and Bestseller Lists
5. Accessing Naming Molecular Compounds Pogil Answers Free and Paid eBooks

- Naming Molecular Compounds Pogil Answers Public Domain eBooks
- Naming Molecular Compounds Pogil Answers eBook Subscription Services
- Naming Molecular Compounds Pogil Answers Budget-Friendly Options
- 6. Navigating Naming Molecular Compounds Pogil Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Naming Molecular Compounds Pogil Answers Compatibility with Devices
 - Naming Molecular Compounds Pogil Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Naming Molecular Compounds Pogil Answers
 - Highlighting and Note-Taking Naming Molecular Compounds Pogil Answers
 - Interactive Elements Naming Molecular Compounds Pogil Answers
- 8. Staying Engaged with Naming Molecular Compounds Pogil Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Naming Molecular Compounds Pogil Answers
- 9. Balancing eBooks and Physical Books Naming Molecular Compounds Pogil Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Naming Molecular Compounds Pogil Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Naming Molecular Compounds Pogil Answers
 - Setting Reading Goals Naming Molecular Compounds Pogil Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Naming Molecular Compounds Pogil Answers
 - Fact-Checking eBook Content of Naming Molecular Compounds Pogil Answers
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development

- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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