

# NOMENCLATURE PACKET

Worksheet I: Binary Ionic Compounds (representative metals) - metals from groups 1A, 2A, and 3A (1, 2, and 13) have constant charges as ions and do NOT get Roman Numerals in their names

1. Name the following ionic compounds:

- |                            |                    |
|----------------------------|--------------------|
| a. $\text{Al}_2\text{O}_3$ | Aluminum oxide     |
| b. $\text{Cs}_2\text{O}$   | Cesium oxide       |
| c. $\text{Rb}_3\text{N}$   | Rubidium nitride   |
| d. $\text{Ca}_3\text{N}_2$ | Calcium nitride    |
| e. $\text{SrSe}$           | Strontium selenide |
| f. $\text{Cs}_2\text{S}$   | Cesium sulfide     |
| g. $\text{Al}_2\text{S}_3$ | Aluminum sulfide   |
| h. $\text{LiBr}$           | Lithium bromide    |
| i. $\text{Mg}_3\text{N}_2$ | Magnesium nitride  |
| j. $\text{CaF}_2$          | Calcium fluoride   |

2. Write the chemical formula for the following ionic compounds:

- |                                |                  |                  |                         |
|--------------------------------|------------------|------------------|-------------------------|
| a. Barium nitride              | $\text{Ba}^{+2}$ | $\text{N}^{-3}$  | $\text{Ba}_3\text{N}_2$ |
| b. Indium Fluoride             | $\text{In}^{+3}$ | $\text{F}^{-1}$  | $\text{InF}_3$          |
| c. Calcium oxide               | $\text{Ca}^{+2}$ | $\text{O}^{-2}$  | $\text{CaO}$            |
| d. Sodium nitride              | $\text{Na}^{+1}$ | $\text{N}^{-3}$  | $\text{Na}_3\text{N}$   |
| e. Magnesium chloride          | $\text{Mg}^{+2}$ | $\text{Cl}^{-1}$ | $\text{MgCl}_2$         |
| f. Potassium oxide             | $\text{K}^{+1}$  | $\text{O}^{-2}$  | $\text{K}_2\text{O}$    |
| g. Magnesium oxide             | $\text{Mg}^{+2}$ | $\text{O}^{-2}$  | $\text{MgO}$            |
| h. Potassium sulfide           | $\text{K}^{+1}$  | $\text{S}^{-2}$  | $\text{K}_2\text{S}$    |
| i. Lithium nitride             | $\text{Li}^{+}$  | $\text{N}^{-3}$  | $\text{Li}_3\text{N}$   |
| j. Strontium fluoride          | $\text{Sr}^{+2}$ | $\text{F}^{-1}$  | $\text{SrF}_2$          |
| k. Aluminum sulfide            | $\text{Al}^{+3}$ | $\text{S}^{-2}$  | $\text{Al}_2\text{S}_3$ |
| l. Duplicate problem - removed |                  |                  |                         |

# Names And Formulas Worksheet 1 Answers

**Ying Liu**



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## **Table of Contents Names And Formulas Worksheet 1 Answers**

1. Understanding the eBook Names And Formulas Worksheet 1 Answers
  - The Rise of Digital Reading Names And Formulas Worksheet 1 Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Names And Formulas Worksheet 1 Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Names And Formulas Worksheet 1 Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Names And Formulas Worksheet 1 Answers

- Personalized Recommendations
  - Names And Formulas Worksheet 1 Answers User Reviews and Ratings
  - Names And Formulas Worksheet 1 Answers and Bestseller Lists
5. Accessing Names And Formulas Worksheet 1 Answers Free and Paid eBooks
- Names And Formulas Worksheet 1 Answers Public Domain eBooks
  - Names And Formulas Worksheet 1 Answers eBook Subscription Services
  - Names And Formulas Worksheet 1 Answers Budget-Friendly Options
6. Navigating Names And Formulas Worksheet 1 Answers eBook Formats
- ePub, PDF, MOBI, and More
  - Names And Formulas Worksheet 1 Answers Compatibility with Devices
  - Names And Formulas Worksheet 1 Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
- Adjustable Fonts and Text Sizes of Names And Formulas Worksheet 1 Answers
  - Highlighting and Note-Taking Names And Formulas Worksheet 1 Answers
  - Interactive Elements Names And Formulas Worksheet 1 Answers
8. Staying Engaged with Names And Formulas Worksheet 1 Answers
- Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Names And Formulas Worksheet 1 Answers
9. Balancing eBooks and Physical Books Names And Formulas Worksheet 1 Answers
- Benefits of a Digital Library
  - Creating a Diverse Reading Collection Names And Formulas Worksheet 1 Answers
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Names And Formulas Worksheet 1 Answers
- Setting Reading Goals Names And Formulas Worksheet 1 Answers
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Names And Formulas Worksheet 1 Answers

- Fact-Checking eBook Content of Names And Formulas Worksheet 1 Answers
- Distinguishing Credible Sources

### 13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

### 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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